SUSHRUT

Mentors Eduserv®

Test Booklet Code-70

Do **not open** this Test **Booklet until you are asked to do so.**

Important instructions:

1. The Answer Sheet is inside this Test Booklet. When you are directed to open the Test Booklet, take out the Answer Sheet and fill in the particulars on ORIGINAL Copy carefully with blue/black ball point pen only.
2. The test is of **3 hours 15 minutes** duration and the Test Booklet contains 180 multiple-choice questions (four options with a single correct answer) from **Physics, Chemistry and Biology (Botany and Zoology).**
3. Wherever the symbols/constants are not mentioned, they are to be considered as per their standard meaning/value.
4. Each question carries **4 marks**. For each correct response, the candidate will get **4 marks**. For each incorrect response, one mark will be deducted from the total scores. **The maximum marks are 720.**
5. Use **Blue/Black Ball Point Pen only** for writing particulars on this page/marking responses on Answer Sheet.
6. Rough work is to be done in the space provided for this purpose in the Test Booklet only.
7. On completion of the test, the candidate **must handover the Answer Sheet (ORIGINAL and OFFICE Copy) to the Invigilator** before leaving the Room/Hall. The candidates are allowed to take away this Test Booklet with them.
8. The **CODE for this Booklet is "70". Make sure to enter this code in the OMR answer sheet.**
9. The candidates should ensure that the Answer Sheet is not folded. Do not make any stray marks on the Answer Sheet. Do not write your Roll No. anywhere else except in the specified space in the Test Booklet/Answer Sheet.
10. Use of white fluid for correction is NOT permissible on the Answer Sheet.
11. Each candidate must show on demand his/her Admit Card to the Invigilator.
12. No candidate, without special permission of the centre Superintendent or Invigilator, would leave his/her seat.
13. The candidates should not leave the Examination Hall without handing over their Answer Sheet to the Invigilator on duty and sign (with time) the Attendance Sheet twice. **Cases, where a candidate has not signed the Attendance Sheet a second time, will be deemed not to have handed over the Answer Sheet and dealt with as an Unfair Means case.**
14. Use of Electronic/Manual Calculator is prohibited.
15. The candidates are governed by all Rules and Regulations of the examination with regard to their conduct in the Examination Room/Hall. All cases of unfair means will be dealt with as per the Rules and Regulations of this examination along with Public Examinations (Prevention of unfair means act 2024).
16. **No part of the Test Booklet and Answer Sheet shall be detached under any circumstances.**
17. The candidates will write the Correct Test Booklet Code as given in the Test Booklet/Answer Sheet in the Attendance Sheet.
18. If a candidate marks more than one answers for a question in the **OMR sheet**, it will be treated as incorrect and negative marking will be applicable.

Name of the Candidate (in Capitals): _____

Roll Number : in figures _____

: in words _____

Centre of Examination (in Capitals) : _____

Candidate's Signature : _____ Invigilator's Signature : _____

Facsimile signature stamp of Centre Superintendent : _____

RE-NEET UG-2026 (EXAMINATION)

(Held on 21st JUNE 2026)

TEST PAPER WITH ANSWER KEY & DETAILS SOLUTION

PHYSICS

[Q.1] A photon and an electron, each of 20 eV energy, move in free space. The ratio of linear momentum of electron P_e to that of photon P_{ph}, $\frac{P_e}{P_{ph}}$ is :

(Take speed of light = 3×10^8 ms^{-1} charge of electron = -1.6×10^{-19} C and mass of electron = 9×10^{-31} kg)

(1) 275 (2) $\frac{2}{450}$ (3) $\frac{1}{250}$ (4) 225

[ANS] **4**

[SOLN] Electron = 20 eV Ephoton = 20 eV

$P=\sqrt{2mKE}$ $P_{ph}=\frac{E}{C}=\frac{20\times1.6\times10^{-19}}{3\times10^8}$

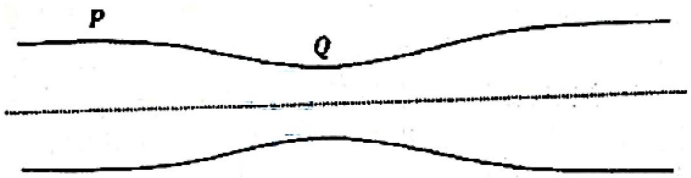
$=\sqrt{2\times9.1\times10^{-31}\times20\times1.6\times10^{-19}}$

$$\frac{P_e}{P_{pn}}=\frac{\sqrt{2\times9.1\times10^{-31}\times20\times1.6\times10^{-19}}}{\frac{20\times1.6\times10^{-19}}{3\times10^8}}$$

= 225 Ans.

[Q.2] Water flows in a streamline motion through a horizontal pipe of circular cross-section as shown in the figure. The pressure difference of water between P and Q is 15 Nm^{-2}. The area of cross-section at P and Q are 40 cm^2 and 20 cm^2, respectively. The rate of flow of water through the pipe, in cm^{3}s^{-1}, is :

[Take density of water = 1000 kg m^{-3}]



(1) 400 (2) 100 (3) 200 (4) 300

[ANS] **1**

[SOLN] $A_1V_1 = A_2V_2$

$40 \times V_1 = 20 \times V_2$

$2V_1 = V_2$

$P_1 + \frac{1}{2}\rho V_1^2 = P_2 + \frac{1}{2}\rho V_2^2$

$P_1 - P_2 = \frac{1}{2}\rho\left(V_2^2 - V_1^2\right)$

$(15) \times 2 = 1000\left(3V_1^2\right)$

$\frac{30}{3000} = V_1^2$

$V_1 = \frac{1}{10}$ m/s

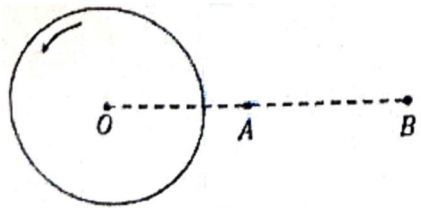
$V_1 = 10$ cm/sec

$Q = A_1V_1$

$= 40 \times 10$

$= 400 cm^3$ / sec

[Q.3] A thin horizontal disc is rotating about a vertical axis passing through its fixed centre O. Its angular momentum is L_A and L_B computed about points A and B, respectively, with OB = $2 \times OA$. The value of $\frac{L_A}{L_B}$ is :



(1) 2 (2) $\frac{1}{4}$ (3) $\frac{1}{2}$ (4) 1

[ANS] **4**

[SOLN] $\vec{L} = \vec{L} + m\left(8 \times V_{cm}\right)$

$\vec{V}_{cm} = 0$

$\therefore \quad \frac{L_A}{L_B} = 1$

[Q.4] Consider a long solenoid of length *l* and radius r. If n is the number of turns per unit length and μ_0 is the permeability of free space, the inductance of the solenoid is:

(1) $2\mu_0\pi n^2 r^2 l$ (2) $\mu_0\pi n^2 r^2 l$ (3) $\mu_0 n^2 r^2 l$ (4) $(\mu_0 / 2\pi) n^2 r^2 l$

[ANS] **2**

[SOLN] $\mu_0 n^2 v$

Where v = volume = $\pi r^2 l$

So

$\mu_0 n^2 (\pi r^2 l)$

[Q.5] The temperature of a metallic sphere of radius R is increased by a small amount ΔT. If the linear coefficient of thermal expansion of the metal is α, the approximate increase in the volume of the sphere is :

(1) $6\pi R^3 \alpha \Delta T$ (2) $2\pi R^3 \alpha \Delta T$ (3) $3\pi R^3 \alpha \Delta T$ (4) $4\pi R^3 \alpha \Delta T$

[ANS] **4**

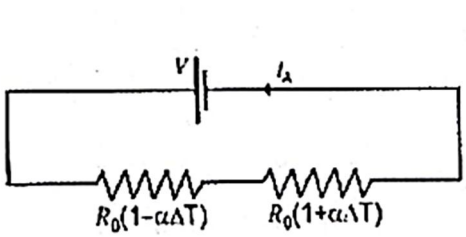
[SOLN] dv = $\gamma v.dt$

or $\Delta v = \gamma v.\Delta T$

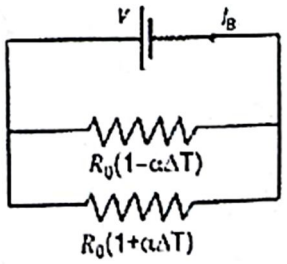
so ,

= $4\pi R^3 \alpha \Delta T$

[Q.6] Consider two circuits, (A) and (B), each having two resistors. One of them has a positive temperature coefficient of resistance, $+\alpha$, while the other one has a negative temperature of coefficient, $-\alpha$, as shown in the figure. The current through these circuits are denote by I_A and I_B. At initial temperature, the resistance of the two resistors is R_0. As the temperature is increased, the correct option that describes the variation of current in these circuits is:



(A)



(B)

(1) both I_A and I_B remain constant (2) I_A remains constant while I_B increases

(3) I_A decreases while I_B increases (4) I_A increases while I_B decreases

[ANS] **2**

[SOLN] for current I_A

$R_{series} = R_1 + R_2$

$= R_0(1-\alpha\Delta T) + R_O(1+\alpha\Delta T)$

$= \boxed{2R_O} \rightarrow cons\tan t$

So ;

$I_A \rightarrow$ Constant

For current I_B

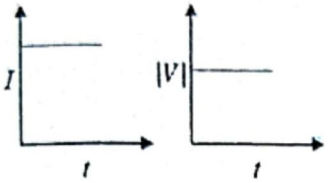
$$R_{eq} = \frac{R_1R_2}{R_1+R_2} = R_B$$

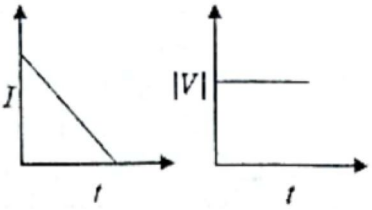
$$= \frac{R_O(1-\alpha\Delta T)R_O(1+\alpha\Delta T)}{2R_O}$$

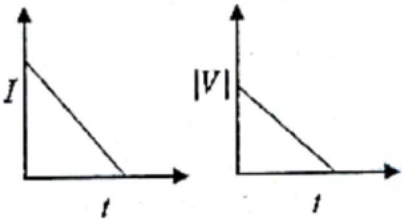
$$= \frac{R_O^2}{2R_O}\left(1^2 - \alpha^2\Delta T^2\right)$$

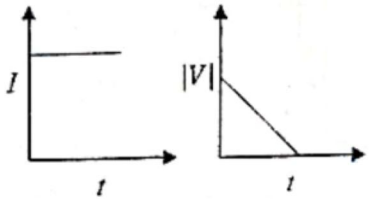
So R_B decreases so I_B increases

[Q.7] A beam of light falls on a metal surface such that photo-electrons are generated. If power of the light source starts to decrease linearly with time t, then variation of the photocurrent I and magnitude of the stopping potential $|V|$ with time is best represented by :

(1) 

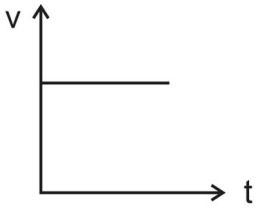
(2) 

(3) 

(4) 

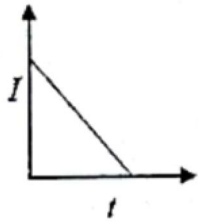
[ANS] **2**

[SOLN] Stopping potential independent of time

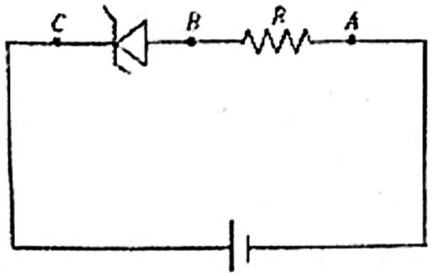


Power decreases linearly with time

So,



[Q.8] An ideal Zener diode with breakdown voltage of –3V is reverse biased with a negative input voltage $V_i = -5V$. The magnitude of voltage difference between points B and A is :



(1) 0 V (2) 3 V (3) 2 V (4) 1 V

[ANS] **3**

[SOLN] breakdown = (–3V)

Vinput = –5V

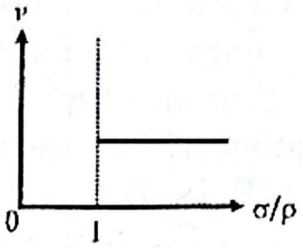
A/q

$5V = 3V + V_{(BA)}$

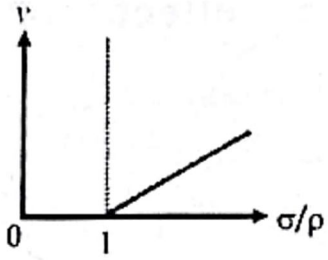
$V_{BA} = 2$ Vol ^{T}

[Q.9] In the measurement of viscosity of liquids using terminal velocity experiment, spherical balls of same radius but having different densities are used. The variation of the terminal velocity (v) with the ratio of density of spherical ball (σ) to density of the liquid (ρ), is best represented by

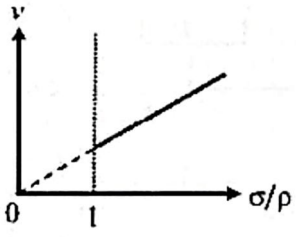
(1)



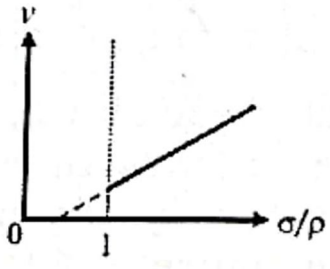
(2)



(3)

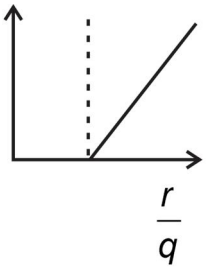


(4)



[ANS] 2

[SOLN] $V_T \propto \frac{2}{9}\frac{r^2(FQ)\cdot g}{n}$



[Q.10] Two planets P_1 and P_2 with equal mass have radii R_1 and R_2, respectively. Where $R_2 = \frac{R_1}{R_2}$.

The escape speeds of P_1 and P_2 are ν_1 and ν_2, respectively. Then $\frac{\nu_2}{\nu_1}$

(1) 2 (2) $\frac{1}{\sqrt{2}}$ (3) 1 (4) $\sqrt{2}$

[ANS] **4**

[SOLN] escape speed = $\sqrt{\frac{2GM}{R}}$

$\frac{V_2}{V_1} = \sqrt{\frac{R_1}{R_2}} = \sqrt{2}$ ……(1)

[Q.11] An ac voltage V = 220sin(2×10^3t) Volt is applied to a series LCR circuit. Then the current amplitude in this circuit is:

(Given: L = 10mH , $C = 25\mu F$, R = 100Ω)

(1) 22.0 A (2) 2.2 A (3) 5.5A (4) 11.0 A

[ANS] **2**

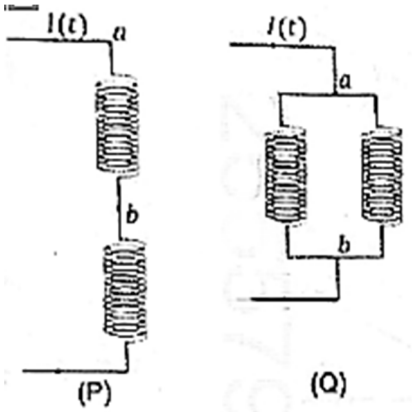
[SOLN] $V = 200\sin(2000t)$

$V = \frac{V_0}{\sqrt{(X_L - X_C^2) + R^2}} = \frac{200}{\sqrt{(20-20)^2 + 100^2}} = \frac{220}{100} = 2.2\,A$

$X_C = \frac{1}{2\pi fC} = \frac{1}{2000\times25\times10^{-6}}$

$X_L = \omega L = 2000\times10\times10^{-5} = 20\,\Omega$

[Q.12] Two identical inductors are connected in two different configurations P and Q, where a time varying current I(t) is flowing, as shown in the figure. The induced emf between points a and b for configuration P is E_p and that for configuration Q is E_Q. The ratio E_p/E_Q is:
[Neglect the effect of mutual inductance.]



(1) 2 (2) 1/4 (3) 1/2 (4) 1

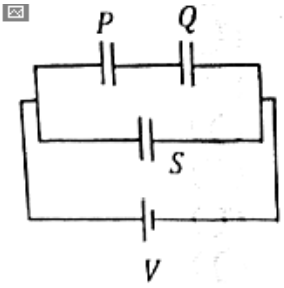
[ANS] **1**

[SOLN] $V_{ab} = E_P = \frac{LdI}{dt}$

$V_{ab} = E_\phi = \frac{1}{2}\frac{d\tau}{dt}$

$\frac{E_P}{E_\phi} = \frac{\frac{LdI}{dt}}{\frac{1}{2}\left(\frac{dI}{dt}\right)} = 2$

[Q.13] Three identical capacitors P, Q and S, each of the capacitance C, are connected to a battery of voltage V, as shown in the figure. If the energy stored in the capacitor P and total energy stored in the system are U_P and U_T respectively, then the ratio is: U_P / U_T is :



(1) 1/6 (2) 2/3 (3)1/3 (4) 1/2

[ANS] **1**

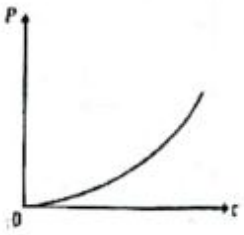
[SOLN] $C_{eq} = \frac{C}{2} + C = \frac{3C}{2}$

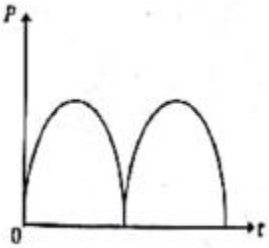
$U_T = \frac{1}{2} C_{eq} V^2 = \frac{1}{2}\left(\frac{3C}{2}\right) V^2 \quad = \frac{3CV^2}{4}$

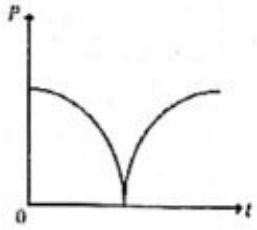
$U_P = \frac{1}{2} C \left(\frac{V}{2}\right)^2 = \frac{CV^2}{8}$

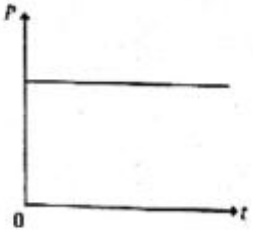
$\frac{U_P}{U_T} = \frac{CV^2}{8} \times \frac{4}{3CV^2} = \frac{1}{6}$

[Q.14] A conducting loop of finite resistance lies on the x - y plane. There is a constant magnetic field in the z direction. The area of the loop varies with time t, as $A = A_0(1+\sin t)$ in appropriate units. The figure that correctly indicates the qualitative behaviour of the power P dissipated in the loop as a function of time is:

(1) 

(2) 

(3) 

(4) 

[ANS] **3**

[SOLN] $\phi = BA = BA_0(1+\sin t)$

$\varepsilon = \frac{-d\phi}{dt} = -BA_0 \cos t$

$P = \frac{\varepsilon^2}{R} = \frac{B^2 A_0^2 \cos^2 t}{R}$

at $t = 0, \quad \cos t = 1$

$P = P_{max} = \frac{B^2 A_0^2}{R}$

at $t = \pi/2, \quad \cos t = 0$

$P = 0$

[Q.15] In an adiabatic expansion, the temperature of one mole of an ideal monatomic gas $(\gamma = 5/3)$ decreases from 60K to 50K. The work done by the gas in the process is:

(Take the universal gas constant as $R = 8.3\ Jmol^{-1}\ K^{-1}$)

(1) 166 J (2) 41.5 J (3) 83 J (4) 124.5J

[ANS] **4**

[SOLN] For adiabatic Forces,

$Q = 0$

$Q = \Delta U + W$

$\Rightarrow W = -\Delta U$

$\Delta U = nC_v\Delta T = n \times \frac{3}{2}R\Delta T$

$\Delta U = \delta \times \frac{3}{2} \times 8.3 \times (50 - 60) = -124.5\,J$

$W = -\Delta U = 124.5\,J$

[Q.16] Consider a particle moving along a straight line, whose position as a function of time is given by $s(t) = \alpha t^2 - \beta t + \gamma$, where $\alpha = 1 ms^{-2}, \beta = 6 ms^{-1}$ and $\gamma = 5\,m$. The average speed of the particle, in ms^{-1} from t = 0 to t = 6 is

(1) 0 (2) 12 (3) 6 (4) 3

[ANS] **4**

[SOLN] $S = \alpha t^2 - \beta t + \gamma$

$S = t^2 - 6t + 5$

$S(0) = 5m$

$S(6) = 6^2 - 6 \times 6 + 5 = 5m$

$S(3) = 3^2 - 6 \times 3 + 5 = -4m$

Total distance $= 9 + 9 = 18m$

$V_{avg} = \frac{18}{6} = 3\,m/s$

$V = 2t - 6$

at $t = 3\,sec,\ V = 0$

[Q.17] The following table presents the part of the electromagnetic spectrum and their corresponding major applciations.

Part of the electromagnetic spectrum		Applications	
P	Microwave	I	For purifying the water
Q	UV rays	II	For warming the food
R	Gamma rays	III	For AM and FM communication systems
S	Radio wave	IV	For treating the Cancer cells

The correct option is:

(1) P-II, Q-IV, R-III, S-I (2) P-I, Q-II, R-III, S-IV

(3) P-I, Q-IV, R-II, S-III (4) P-II, Q-I, R-IV, S-III

[ANS] **4**

[SOLN] Conceptual

[Q.18] An ideal gas is made of polyatomic molecules. Each of the molecules has three translational, three rotational and f number of vibrational modes. If the ratio of heat capacities C_p / C_v of the gas is 8/7, then the value of f is:

(1) 1 (2) 4 (3) 3 (4) 2

[ANS] **2**

[SOLN] gamma = $1+\frac{2}{f}=\frac{8}{7}$

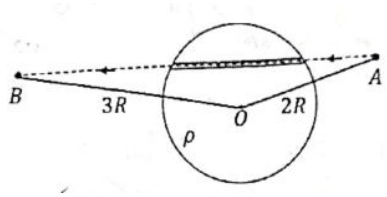
F = 14

3 + 3 + 2V = 14

V = 4

[Q.19] A unit positive point charge is taken slowly through an infinitesimally thin tube that is inside a charged dielectric sphere of radius R, having uniform positive charge density as shown in the figure. The initial and final positions of the charge are marked by A and B at distances 2R and 3R respectively, from the centre of the sphere. In this process, the magnitude of the total work done on the point charge is $\frac{\rho R^2}{n\in_0}$. The value of n is:

($\in_0$ is the permittivity of vacuum)

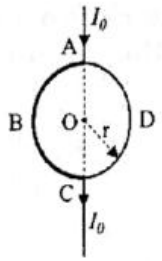


(1) 18 (2) 2 (3) 6 (4) 9

[ANS] **1**

[SOLN] **work done =** $q\Delta V = q\left(V_B - V_A\right) = \dfrac{\rho R^2}{18\,\epsilon_0}$

[Q.20] A current I_o flows through a metallic circular loop of radius r as shown in the figure. Resistance of the segment ABC is half that of ADC. Magnitude of magnetic field at the center O of the loop is:



(1) $\dfrac{\mu_0 I_0}{2\pi r}$ (2) $\dfrac{\mu_0 I_0}{12r}$ (3) $\dfrac{\mu_0 I_0}{4r}$ (4) $\dfrac{\mu_0 I_0}{2r}$

[ANS] **2**

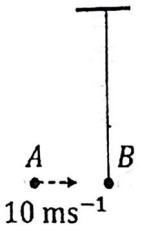
[SOLN] $i_0 = i_1 + i_2$

$i_1 = \dfrac{2I_0}{3}(\text{ABC})$

$i_2 = \dfrac{I_0}{3}(\text{ADC})$

$B = \dfrac{\mu_0 i_1}{4r}\odot + \dfrac{\mu_0 i_2}{4r}\otimes = \dfrac{\mu_0 I_0}{12r}$

[Q.21] Bob B of mass m at rest is hanging vertically from the ceiling via a massless string of length 10 m, as shown in the figure. Point mass A of mass m travelling horizontally with speed $10\,ms^{-1}$ hits bob B elastically. The bob B rises h meter after the collision. Taking the acceleration due to gravity $g = 10\,ms^{-2}$ and neglecting the size of the bob, the value of h is :



(1) 2.5 (2) 8 (3) 7 (4) 5

[ANS] **4**

[SOLN] $\dfrac{1}{2}mv^2 = mgh$

$v = 5\,m/s$

[Q.22] An electromagnetic wave travelling in a lossless dielectric medium having a dielectric constant, $\epsilon_r = 9$, has the electric field, $E_x = E_0 \sin\left(kz - 2\pi \times 10^6 t\right) Vm^{-1}$ where E_0 is the amplitude and k is the wave vector. Among the following options, the incorrect choice is :

(1) The direction of propagation of the electromagnetic wave is along $+z$

(2) The speed of the electromagnetic wave inside the medium is $10^8\, ms^{-1}$

(3) The wavelength of the electromagnetic wave inside the medium is 300 m

(4) The magnetic field is given by the relation $B_y = \frac{B_0}{\nu}\sin\left(kz - 2\pi \times 10^6 t\right)$ where ν is the speed of the electromagnetic wave inside the medium

[ANS] **Bonus (3 & 4)**

[SOLN] $V = \frac{C}{n} = \frac{C}{\sqrt{\varepsilon r}} = \frac{C}{3} = 10^8\, m/s$

$f = 10^6\, Hz$

$\lambda = 100\, m$

(3) & (4) one incorrect

[Q.23] A particle of mass M moves along a horizontal x axis from $x = 0$ to $x = L$. The coefficient of kinetic friction varies as a function of x as $\mu_k(x) = \mu_0 - \alpha x$, where μ_0, α are constants of appropriate dimension, so that $\mu_k(L) = 0$. The total work done by the frictional force during the motion is $n\mu_0 MgL$, where g is the acceleration due to gravity. The value of n is :

(1) $\frac{1}{2}$ (2) 3 (3) 1 (4) $\frac{1}{3}$

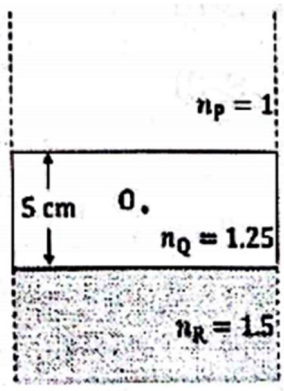
[ANS] **1**

[SOLN] $dw = (\mu_0 - \alpha x) Mg \cdot dx$

$dw = \alpha(L - x) Mg \cdot dx$

$w = Mg\alpha \int_0^L (L - x) \cdot dx = \frac{Mg\alpha L^2}{2} = \frac{Mg\mu_0 L}{2}$

[Q.24] Consider three media P, Q and R with refractive indices 1, 1.25, and 1.5, respectively. The medium Q having a thickness of 5 cm is placed between extended media P and R as shown in the figure. An object O is placed at the center of medium Q. If viewed from medium P near the normal direction, the apparent depth of O is h_1. For similar observation is h_2. The value of $|h_1 - h_2|$, in cm, is :



(1) 3 (2) 0 (3) 1 (4) 2

[ANS] **3**

[SOLN] $\frac{h_1}{2.5} = \frac{1}{1.25} \Rightarrow h_1 = 2\,\text{cm}$

$\frac{h_2}{2.5} = \frac{1.5}{2.5} \Rightarrow h_1 = 3\,\text{cm}$

[Q.25] Consider a fixed uniformly charged insulating sphere with radius R and total charge $+Q$. A point charge $-q(q << Q)$ with mass m is released from rest at a distance of 3R from the centre of the charged sphere. When the point charge reaches the surface of the sphere, its speed is :

($\in_0$ is the permittivity of vacuum, neglect gravitational forces).

(1) $\sqrt{\frac{Qq}{4\pi \in_0 mR}}$ (2) $\sqrt{\frac{3Qq}{4\pi \in_0 mR}}$ (3) $\sqrt{\frac{2Qq}{3\pi \in_0 mR}}$ (4) $\sqrt{\frac{Qq}{3\pi \in_0 mR}}$

[ANS] **4**

[SOLN] increase in KE = dec is potential energy

$\frac{1}{2}mv^2 = q\left(\frac{K\theta}{R} - \frac{K\theta}{3R}\right)$

$v = \sqrt{\frac{Qq}{3\pi \in_0 mR}}$

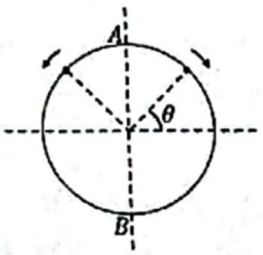
[Q.26] A car travels on a circular racetrack of radius 50 m, which is banked at an angle θ. If the car travels at a speed $10\,ms^{-1}$, then the wear and tear on its types is minimum. Taking the acceleration due to gravity to be $10\,ms^{-2}$, the value of θ is :

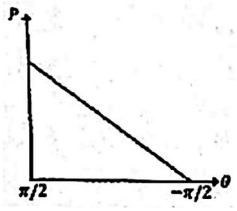
(1) $\tan^{-1}\left(2\sqrt{3}\right)$ (2) $\tan^{-1}\left(\frac{1}{5}\right)$ (3) $\tan^{-1}\left(\frac{2}{5}\right)$ (4) $\tan^{-1}\left(\sqrt{3}/2\right)$

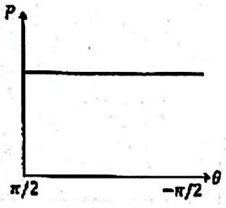
[ANS] **2**

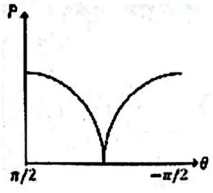
[SOLN] $\tan\theta = \frac{v^2}{rg} = \frac{1}{5}$

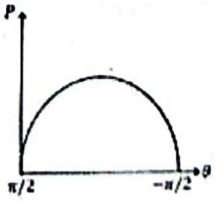
[Q.27] A frictionless circular wire of unit radius is fixed on the horizontal plane. Two point particles of unit mass start moving simultaneously from point A $\left(\theta = \frac{\pi}{2}\right)$ with identical uniform angular speeds in opposite directions, and meet again at point B $\left(\theta = -\frac{\pi}{2}\right)$. During this time, which of the following figures schematically represent the magnitude of the total linear momentum P of the system, as a function of θ?



(1) 

(2) 

(3) 

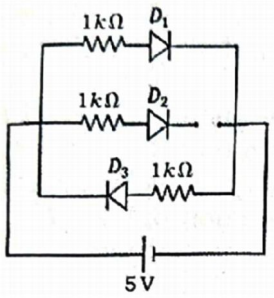
(4) 

[ANS] **4**

[SOLN] Initial $\left(\theta = \frac{\pi}{2}\right) \rightarrow P_{net} = 0$

$\left(\theta = \frac{-\pi}{2}\right) \rightarrow P_{net} = 0$

[Q.28] Three identical p-n junction diodes D_1, D_2 and D_3 are connected across a battery as shown in the figure. If the width of the depletion regions of D_1, D_2 and D_3 are W_1, W_2 and W_3, respectively, then the correct option is :



(1) $W_2 > W_1 = W_3$ (2) $W_1 > W_2 > W_3$ (3) $W_3 = W_1 > W_2$ (4) $W_3 > W_2 > W_1$

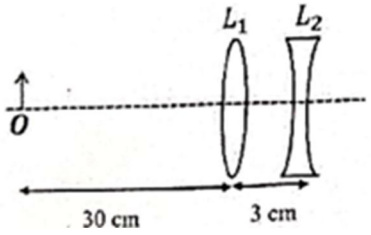
[ANS] **4**

[SOLN] Forward bias $\rightarrow$ width of depletion region

Reverse $\rightarrow$ decreases

$\therefore W_3 > W_2 > W_1$

[Q.29] The lens combination as shown in the figure, consists of two lenses, L_1, and L_2, of the focal lengths +10 cm and —10 cm, respectively. The position of the image formed is:



(1) 60 cm to the right of the concave lens (2) 20 cm to the left of the concave lens

(3) 60 cm to the left of the concave lens (4) 30 cm to the right of the concave lens

[ANS] **3**

[SOLN] $v = \dfrac{uf}{u+f}$

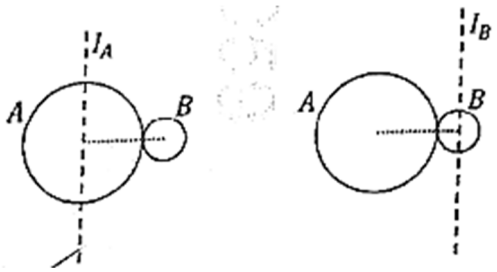
For lens 1

$v = \dfrac{(-30)(10)}{-30+10} = 15\,\text{cm}$

For lens 2

$v = \dfrac{(12)(-10)}{12-10} = \dfrac{-120}{2} = -60\ \text{cm}$

[Q.30] A solid sphere A of radius R and Mass M is attached at a point to a Smaller solid sphere B of radius r < R and mass m < M. Assume that the line joining their centers lies along the horizontal. The moment of inertia of the System, calculated about a vertical axis passing through the Centre of A is I_A, and that calculated about a vertical axis Passing through the centre of B is I_B. Th, difference $I_A - I_B$ is:



(1) 0 (2) $(M-m)(R+r)^2$ (3) $(m-M)(R+r)^2$ (4) $(m-M)(R-r)^2$

[ANS] **3**

[SOLN] $I_A = \frac{2}{5}MR^2 + \frac{2}{5}mr^2 + m(R+r)^2$

$I_B = \frac{2}{5}mr^2 + \frac{2}{5}MR^2 + M(R+r)^2$

$\therefore I_A - I_B = (m-M)(R+r)^2$

[Q.31] Consider that an electron is revolving in an excited state of Hydrogen atom with velocity $\sqrt{25.6}\times10^5\,ms^{-1}$. The radius of the orbit is $x\times10^{-9}m$. The value of x is :

[Take the mass of electron to be $9\times10^{-31}kg$,

charge of electron $= -1.6\times10^{-19}C$ and $\frac{1}{4\pi\varepsilon_o} = 9\times10^9\,N\,m^2C^{-2}$]

(1) 1 (2) 4 (3) 3 (4) 2

[ANS] **1**

[SOLN] $\frac{mv^2}{r} = \frac{kq^2}{r^2}$

$\Rightarrow r = \frac{kq^2}{m_e v^2}$

Here

$v = \sqrt{25.6}\times10^5$

So, $v^2 = 25.6\times10^{10}$

So, $r = \frac{9\times10^9\times(1.6\times10^{-19})^2}{9\times10^{-31}\times(25.6\times10^{10})} = 1\times10^{-9}m$

[Q.32] The mean free path of molecules in an ideal gas A is half that of another ideal gas B. The diameter of the spherical molecules of gas A is twice the diameter of the molecules of B. If number densities of the gases A and B are n_A and n_B, respectively, then the correct options is:

(1) $n_A = \frac{1}{2}n_B$ (2) $n_A = n_B$ (3) $n_A = 2n_B$ (4) $n_A = \frac{1}{4}n_B$

[ANS] **1**

[SOLN] $\lambda_A = \frac{1}{\sqrt{2}\pi d_A^2 n_A}$

$\lambda_B = \frac{1}{\sqrt{2}\pi d_B^2 n_B}$

So, $\frac{\lambda_A}{\lambda_B} = \frac{d_B^2 n_B}{d_A^2 . n_A} = \left(\frac{1}{2}\right)^2 \left(\frac{n_B}{n_A}\right)$

$\Rightarrow \frac{1}{2} = \frac{1}{4} \times \frac{n_B}{n_A}$

$= n_A = \frac{1}{2}n_B$

[Q.33] A cylindrical cork of uniform density floats in a liquid of density ρ_1. If the cork is depressed slightly and released, It oscillates harmonically with time period T. If the same cork floats in another liquid of density ρ_2, then the similar oscillation has time period 2T. The value of ρ_2 / ρ_1 is.

(1) 1/4 (2) 4 (3) 2 (4) 1/2

[ANS] **1**

[SOLN] $T = 2\pi\sqrt{\frac{m}{A\rho_1 . g}}$

And $2T = 2\pi\sqrt{\frac{m}{A\rho_2 . g}}$

So, $\frac{T}{2T} = \sqrt{\frac{\rho_2}{\rho_1}}$

$\Rightarrow \frac{1}{4} = \frac{\rho_2}{\rho_1}$

[Q.34] For sound waves, if the number of nodes for the 5^{th} harmonic of an open-ended pipe is n and that for the 9^{th} harmonic of the same pipe with one of its ends closed is m, the ratio $\frac{n}{m}$ is:

(1) 3/5 (2) 5/9 (3) 9/5 (4) 1

[ANS] **4**

[SOLN] Open pipe

N = 5

Closed Pipe

$$M = \frac{(\text{Work of harm}^{m} + 1)}{2}$$

$$= \frac{9+1}{2} = 5$$

So, Ratio $\frac{N}{M} = \frac{5}{5} = 1$

[Q.35] Consider the following nuclear reaction:

$^{238}U \rightarrow ^{234}Th + ^{4}He$

Take masses of ^{238}U, ^{234}Th and ^{4}He as 238.050 u, 234.043 u and 4.003 u, respectively. The Q value for the reaction, in keV, is [Given: 1 u = 931.5 MeV c^{-2}]

(1) 3740 (2) 3726 (3) 3730 (4) 3736

[ANS] **2**

[SOLN] $^{238}U \rightarrow ^{234}Th + ^{4}He + \text{Energy}$

Mass Cousn

238 = 234 + 4

$\Delta m = 238.05 - 234.043 - 4.003 = 0.0046$

So

Energy read = Δmc^2

$= 0.0046 \times 931.5$ MeV.

[Q.36] Which of the following measurements require 'index correction' ?

(1) Measurement of speed of sound using resonance tube

(2) Measurement of resistance of a wire using meter bridge

(3) Measurement of gravitational acceleration using simple pendulum

(4) Measurement of focal length of lenses using optical bench.

[ANS] **4**

[SOLN] Conceptual

Index correction mostly used in those experiments in which measuring instruments does not show accurate reading.

Mostly Common experiments

Vernier Callipers $\rightarrow$ index correction in case of zero error.

Screw Gauge $\rightarrow$ index correction in case of zero error.

Any many more experiment

Here in given option, most accurate & precise is " Measurement of focal length of lens using optical bench."

[Q.37] In a solar system, the time period of revolution of a planet tracing a circular orbit of radius R is proportional to

(1) R^3 (2) $R^{1/2}$ (3) $R^{3/2}$ (4) R^2

[ANS] **3**

[SOLN] $T^2 \alpha R^3$

$T \alpha R^{\frac{3}{2}}$

[Q.38] Consider that σ_s. k_B, b represent Stefan-Boltzmann constant and Wien's displacement law constant, respectively. The dimension of $\sigma_s k_B^{-1} b$ is :

(1) $[L^{-1}T^{-1}K^{-4}]$ (2) $[L^{-1}T^{-1}K^{-2}]$ (3) $[L^{-1}K^{-2}]$ (4) $[L^{-1}T^{-1}K^{-3}]$

[ANS] **2**

[SOLN] using $\varepsilon_0 n$ 1 , 2 , 3

$$\frac{T_s}{k_b . b} = L^{-1} T^{-1} K^{-2}$$

$$\frac{E}{t} = \sigma A . T^4$$

$$\frac{M^1 L^2 T^{-2}}{T^1 L^2 K^4} = \sigma \ -(1)$$

$$E = \frac{3}{2} K_B . T$$

$$\frac{M^1 L^2 T^{-2}}{K^1} = K_B \ -(2)$$

$$\lambda_{MT} = h$$

$$[L . K] = h \ -(3)$$

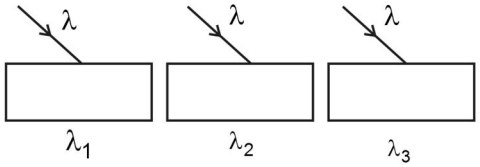
[Q.39] A ray of light with wavelength λ is incident on three different photoelectric cells namely 1, 2 and 3. The threshold wavelength of these photoelectric cells are λ_1, λ_2 and λ_3, respectively and the magnitude of stopping potentials of these cells are V_1, V_2 and V_3, respectively. The relation between λ and threshold wavelengths are $\lambda_1 < \lambda$, $\lambda_2 > \lambda$ and $\lambda_3 >> \lambda$. The correct option is

(1) $V_1 < V_2, V_3 = 0$ (2) $V_1 = 0. \ V_2 < V_3$

(3) $V_1 = 0. \ V_2 > V_3$ (4) $V_1 > V_2. \ V_3 = 0$

[ANS] **2**

[SOLN]



For P.E.E

$Ein > E_{\phi}$

So Here

$V_1 = 0$

$V_1 < V_3$

[Q.40] One main scale division of a Vernier calliper is equal to 1 mm and the number of division on the Vernier scale is 10. When both jaws touch each other, the Verniner scale shifts to the left of zero of the main scale in such a way that 4^{th} Verniner division coincides with a division of the main scale. If this Vernier calliper measures the length of a wire to be 1 cm, the actual length of the wire is

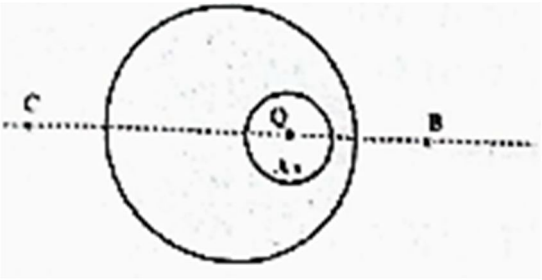
(1) 1.04 cm (2) 0.60 cm (3) 0.96 cm (4) 1.00 cm

[ANS] **Bonus (1)**

[SOLN] Bonus

Option (A)

[Q.41] A point charge Q is placed inside a cavity within a solid isolated conducting sphere. Consider points A, B and C as shown in the figure, where the magnitudes of the electric fields are E_A, E_B and E_C, respectively. The points B and C are a the same distance from the center of the solid sphere. The correct option is.



(1) $E_A \neq 0, E_B < E_C$ (2) $E_A = 0, E_B = E_C$ (3) $E_A \neq 0, E_B = E_C$ (4) $E_A = 0, E_B > E_C$

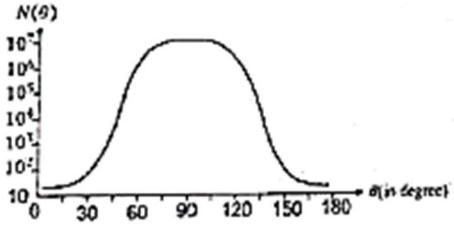
[ANS] **3**

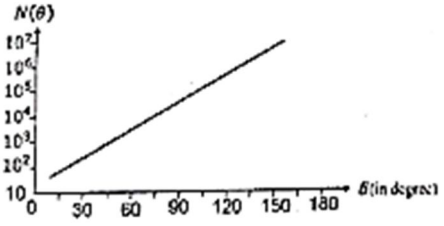
[SOLN] Option – C

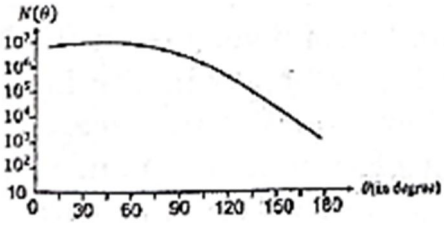
$E_A \neq 0$

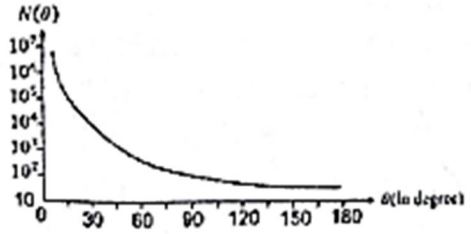
$E_B = E_C$

[Q.42] In Geiger-Marsden experiment, the number of scattered α-particles $N(\theta)$ is plotted as a function of scattering angle θ. Which of the following options represents the correct plot ?

(1) 

(2) 

(3) 

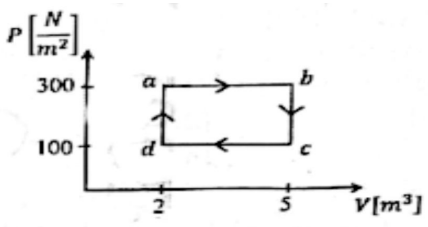
(4) 

[ANS] **4**

[SOLN] $N(\theta) \alpha \dfrac{1}{\text{Sin}^4\left(\dfrac{\theta}{2}\right)}$

Option – 4

[Q.43] One mole of an ideal monatomic gas undergoes a cyclic process as shown in the figure. The total heat supplied to the gas is



(1) 800 J (2) 400 J (3) 500 J (4) 600 J

[ANS] **4**

[SOLN] $\Delta Q = \Delta U + \Delta W$

$\Delta Q = 200 \times 3 = 600 J$

[Q.44] Two infinitely long parallel conducting wires A and B carry currents I and 2I, respectively, in the same direction. The wire A has uniform mass per unit length λ and lies on an insulated floor. The wire B is kept fixed at a height h above the floor. The minimum magnitude of h so that the wire A does not rise from the floor is :

[g is the acceleration due to gravity and μ_0 is permeability of free space.]

(1) $\dfrac{4\mu_0 I^2}{\pi\lambda g}$ (2) $\dfrac{\mu_0 I^2}{2\pi\lambda g}$ (3) $\dfrac{\mu_0 I^2}{\pi\lambda g}$ (4) $\dfrac{2\mu_0 I^2}{\pi\lambda g}$

[ANS] **3**

[SOLN] $\frac{\mu_0 I(2l)}{2\pi h} = \lambda g$

$h = \frac{\mu_0 I^2}{\pi \lambda g}$

[Q.45] Consider a spring-mass simple harmonic oscillator is one dimension. The mass of the particle is m kg and the spring constant k Nm^{-1}. At a given instant, the extension of the spring is x meter and the speed of the particle v ms^{-1}. On the x – v plane, if the graph of v as a function of x is a circle, then the correct option is

(1) $k = \sqrt{m}$ (2) $k = \frac{1}{m}$ (3) $k = m$ (4) $k = m^2$

[ANS] **3**

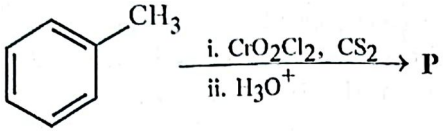
[SOLN] $w = \sqrt{\frac{k}{m}}$

$w = 1$

$K = M$

CHEMISTRY

[Q.46] Consider the following reaction, and choose the correct option.



(1) Compound P is obtained by the hydrogenation of benzoyl chloride with Pd on $BaSO_4$.

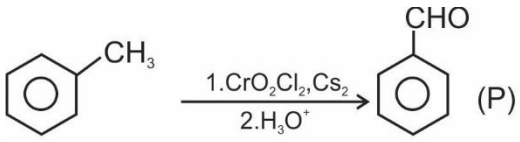
(2) On treating compound P with saturated $NaHCO_3$ solution, brisk effervescence is observed.

(3) Compound P can be prepared by treating benzene with anhydrous $AlCl_3$ and CH_3COCl.

(4) On treatment with bromine water, compound P gives a white precipitate.

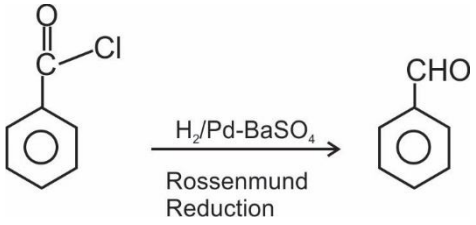
[ANS] **1**

[SOLN] Etard reaction



(Benzaldehyde)

P $\rightarrow$ Obtained from benzoyl chloride reaction with $Pd/BaSO_4$.



[Q.47] The formula of tetraammineaquachloridocobalt(III) chloride is

(1) $[Co(NH_3)_4(H_2O)Cl]Cl_2$ (2) $[Co(NH_3)_4Cl_2] \times H_2O$

(3) $[Co(NH_3)_4]\ Cl_3 \times H_2O$ (4) $[Co(NH_3)_4(H_2O)Cl]Cl$

[ANS] **1**

[SOLN] tetrammineaquachloridocobalt(III) chloride

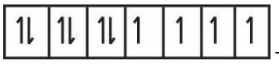
$\left[Co(NH_3)_4(H_2O)Cl\right]Cl_2$

[Q.48] The lanthanide ion having four unpaired electrons is

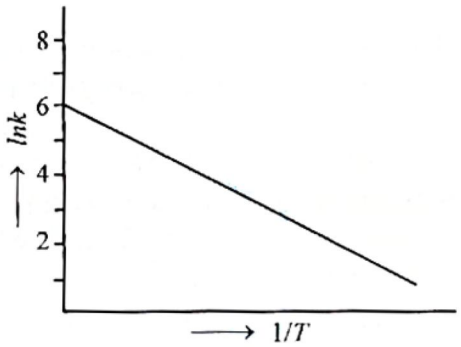
[Given : Atomic numbers of Ce = 58, Nd = 60, Tb = 65 and Ho = 67]

(1) Ho^{3+} (2) Nd^{3+} (3) Ce^{3+} (4) Tb^{3+}

[ANS] **1**

[SOLN] $Ho^{3+} \rightarrow [Xe]4f^{10} \rightarrow$  $\rightarrow$ 4 unpaired electron.

[Q.49] For an elementary chemical reaction, the Arrhenius plot is given below.

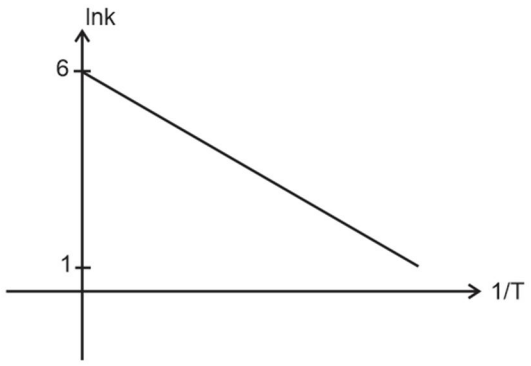


If the energy of activation is 6.64 kJ mol^{-1} and R = 8.3 J K^{-1} mol^{-1}, the temperature at which the rate constant becomes e^2 min^{-1}, is

(1) 250 K (2) 125 K (3) 150 K (4) 200 K

[ANS] **4**

[SOLN]



$E_a = 6.64 kJmol^{-1}$

$R = 8.3\ JK^{-1}mol^{-1}$

$\therefore \quad k = Ae^{-Ea/RT}.$

$\ln A = 6$ therefore $A = e^6$

$k = e^2$

$\therefore\ e^2 = e^6 \times e^{-Ea/RT}$

$$-\frac{E_a}{RT} = -4$$

$$\frac{6640}{8.3 \times T} = 4$$

T = 200 K

[Q.50] The green paramagnetic species formed by heating $KMnO_4$ at 513 K is

(1) KO_2 (2) K_2MnO_4 (3) Mn_3O_4 (4) MnO

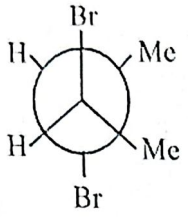
[ANS] **2**

[SOLN] $\underset{\text{purple}}{2KMnO_4} \xrightarrow{\Delta} \underset{\substack{\text{green} \\ (\text{paramagnetic}, n=1)}}{K_2MnO_4} + MnO_2 + O_2$

[Q.51] Given below are two statements:

Statement I:

Trans-But-2-ene upon treatment with Br_2 in CCl_4 givens the following product



In the light of the above statements, choose the **most appropriate** answer from the options given below.

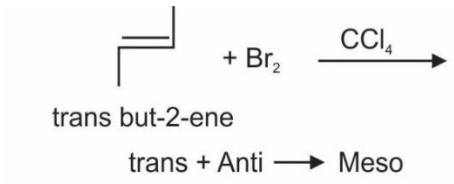
(1) **Statement I** is incorrect but **Statement II** is correct

(2) Both **Statement I** and **Statement II** are correct

(3) Both **Statemtnt I** and **Statement II** are incorrect

(4) **Statement I** is correct but **Statement II** is incorrect

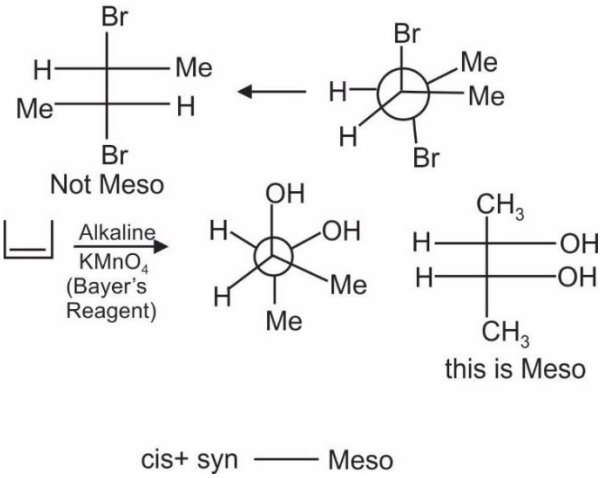
[ANS] **1**



trans but-2-ene

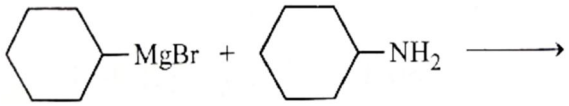
trans + Anti ⟶ Meso

[SOLN]

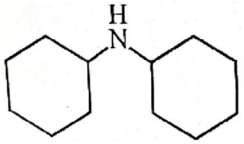


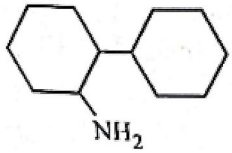
cis+ syn —— Meso

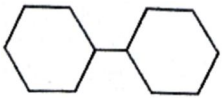
[Q.52] One of the products formed in the following reaction is



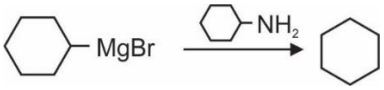
(1) 

(2) 

(3) 

(4) 

[ANS] **1**

[SOLN] 

[Q.53] Given below are two statements:

Statement-I : heating NaCl with concentrated H_2SO_4 and MnO_2 results in oxidation of Mn.

Statement-II : Heating Nal with concentrated H_2SO_4 and MnO_2 results in reduction of Mn.

In light of the above statements, choose the **most appropriate** answer from the options given below:

(1) **Statement-I** is incorrect but **Statement-II** is correct.
(2) Both **Statement-I** and **Statement-II** are correct.
(3) Both **Statement-I** and **Statement-II** are incorrect.
(4) **Statement-I** is correct but **Statement-II** is incorrect.

[ANS] **1**

[SOLN]

Reduction

$$4NaCl + 4H_2SO_4 + MnO_2 \rightarrow 4NaHSO_4 + MnCl_2 + 2H_2O + Cl_2$$

$$2NaI + MnO_2 + 3H_2SO_4 \rightarrow 2NaHSO_4 + MnSO_4 + 2H_2O + I_2$$

Reduction

[Q.54] Among the following options, the correct trend in the electron gain enthalpy is

(1) I > Br > Cl > F (2) F > Cl > Br > I (3) Br > Cl > F > I (4)Cl > F > Br > I

[ANS] **4**

[SOLN] Correct order Cl > F > Br > I

Due to small size of F interelectronic repulsion occurs when electron added so electron affinity of F less than Cl.

[Q.55] Given below are two statements:

Statement-I: $\left[Fe(ox)_3\right]^{3-}$ is chiral.

Statement-II: trans $-\left[Cr(H_2O)_2(OX)_2\right]^-$ is chiral.

(Given: oxH_2 = HOOC–COOH)

In light of the above statements, choose the most appropriate answer from the options given below:

(1) Statement-I is incorrect but Statement-II is correct.

(2) Both Statement-I and Statement-II are correct.

(3) Both Statement-I and Statement-II are incorrect.

(4) Statement-I is correct but Statement-II is incorrect.

[ANS] **4**

[SOLN] $\left[Fe(OX)_3\right]^{3-} \rightarrow$ is chiral because it does not contain any plane of symmetry.

trans-$\left[Cr(H_2O)_2(OX)_2\right]^-$ is not chiral because it contain a plane of symmetry.

[Q.56] The correct statement about peptides and proteins is

(1) In α-helices, the polypeptide chain is twisted into a left-handed screw (helix) through intramolecular hydrogen bonds.

(2) Tertiary structure of proteins has two or more polypeptide subunits.

(3) Only the proteins having a quaternary structure are biologically active.

(4) In β-pleated sheet structures, peptide chains are held together by intermolecular hydrogen bonds.

[ANS] **4**

[SOLN] Biomolecules

[Q.57] Given below are two statements:

Statement-I : Oxidation of p- nitrotoluene with acidic $KMnO_4$ gives an acid that is stronger than benzoic acid.

Statement-II: Reduction of p- nitrotoluene with Sn/HCl followed by neutralization gives an amine that is more basic than aniline.

In light of the above statements, choose the most appropriate answer from the options given below.

(1) Statement-I is incorrect but Statement-II is correct.

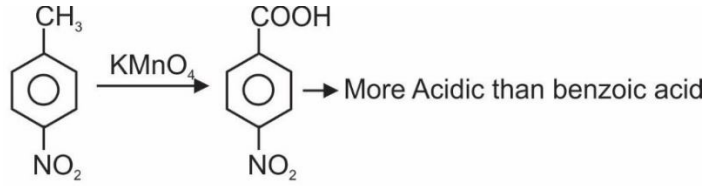
(2) Both Statement-I and Statement-II are correct.

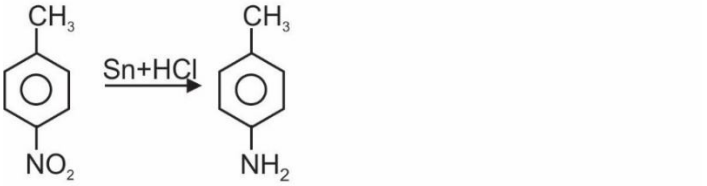
(3) Both Statement-I and Statement-II are incorrect.

(4) Statement-I is correct but Statement-II is incorrect.

[ANS] **2**

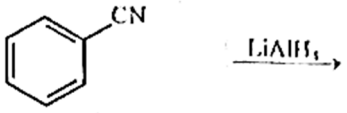
[SOLN]

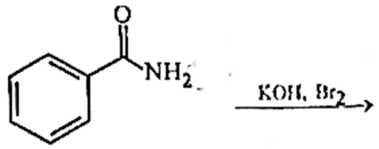


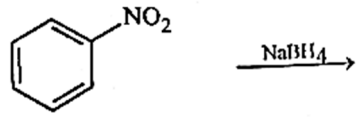


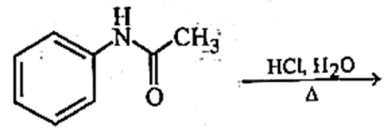
More basic than aniline

[Q.58] Identify the reactions which give aniline as the major product.

A. 

B. 

C. 

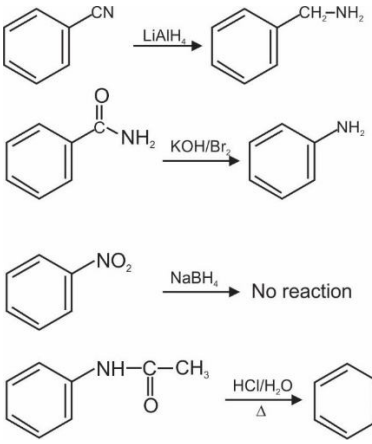
D. 

Choose the correct answer from the options given below.

(1) C and D only (2) A and B only (3) Band D only (4) A and C only

[ANS] **3**

[SOLN]



[Q.59] Two moles of an ideal gas undergo free expansion from 10 L to 100 L at 300 K. The values of ΔS_{system} and $\Delta S_{surroundings}$ are

(R is universal gas constant)

(1) $\Delta S_{system} = 4.606R;;\ \Delta S_{surroundings} = 0$ (2) $\Delta S_{system} = 0; \Delta S_{surroundings} = 0$

(3) $\Delta S_{system} = 4.606; \Delta S_{surroundings} = -4.606R$ (4) $\Delta S_{system} = 0; \Delta S_{surroundings} = 4.606R$

[ANS] 1

[SOLN] Free Expansion (Isothermal)

$$\Delta S_{system} = nR \ln \frac{V_2}{V_1} = 2R \ln \frac{100}{10} = 2R \ln 10 = 2.303x2R \log(10) = 4.606R$$

$$\Delta U = Q + W$$

$W = 0$

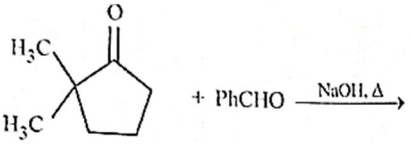
(as Free Expansion)

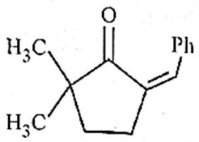
$\Delta U = 0$ (Isothermal)

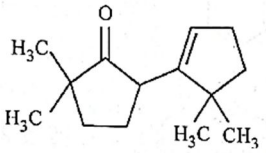
Therefore Q = 0

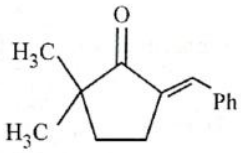
$$\Delta S_{surroundings} = -\frac{Q_{system}}{T_{surrounding}} = -\frac{0}{T_{surrounding}} = 0$$

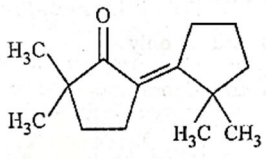
[Q.60] The compound that CANNOT be obtained from the aldol condensation reaction shown below, is



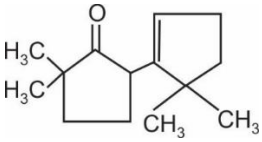
(1) 

(2) 

(3) 

(4) 

[ANS] 2

[SOLN]  Not aldol condensation product.

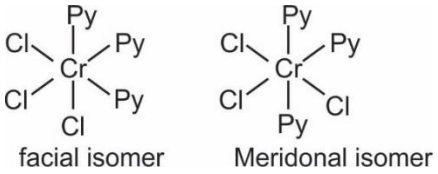
[Q.61] The complex which has facial and meridional isomers is
(Given py pyridine and $H_2N-CH_2-CH_2-NH_2$)

(1) $\left[Ni(en)_2(H_2O)_2\right]^{2+}$ (2) $\left[Cr(py)_3(Cl)_3\right]$

(3) $\left[Cr(H_2O)_6\right]^{3+}$ (4) $\left[Co(NH_3)_4(H_2O)_2\right]^{3+}$

[ANS] **2**

[SOLN] $\left[Cr(Py)_3Cl_3\right]$



[Q.62] The numbers, 17.0145 and 21.0235 were rounded to three figures after the decimal point. The resulting numbers, respectively, are

(1) 17.015 and 21.024 (2) 17.014 and 21.023

(3) 17.015 and 21.023 (4) 17.014 and 21.023

[ANS] **4**

[SOLN] $17.0145\,(\text{Even Rule}) = 17.014$

v

$21.0235\,(\text{Odd Rule}) = 21.024$

[Q.63] The amount of carbon dioxide evolved upon completer combustion of 116 g of n-butane is
(Given : atomic mass in amu $H=1$, $C=12$ and $O=16$)

(1) 362 g (2) 352 g (3) 322 g (4) 176 g

[ANS] **2**

$$C_4H_{10} + \frac{13}{2}O_2 \rightarrow 4CO_2 + 5H_2O$$

$$\text{mol} = \frac{116}{58}$$

$$= 2\,\text{mol}$$

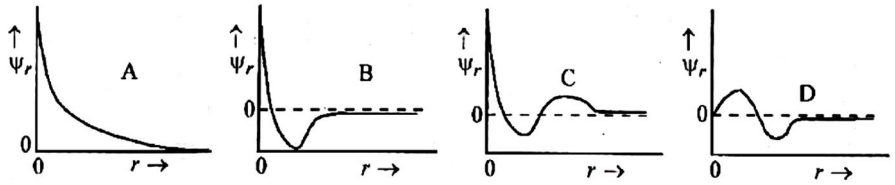
[SOLN] $\frac{\text{mole}_{C_4H_{10}}}{1} = \frac{\text{mole}_{CO_2}}{4}$

$$\frac{2}{1} = \frac{\text{mole}_{CO_2}}{4}$$

$$\text{mole}_{CO_2} = 8$$

$$\text{Mass}_{CO_2} = (8 \times 44)\,\text{gm} = 352\,\text{gm}$$

[Q.64] Consider the following schematic plots of orbital wavefunction (ψ_r) against distance (r) from the nucleus.

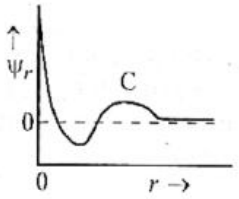


The figure representing two radial nodes in the orbital is

(1) D (2) A (3) B (4) C

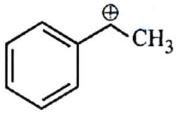
[ANS] **4**

[SOLN]



No. of radial nodes = No. of Cuts on r axis=2

[Q.65] The following carbocation is stabilized by the interaction of the empty p orbital with

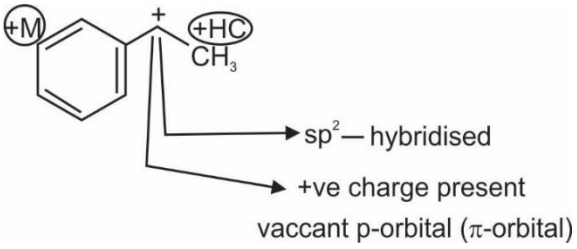


(1) empty σ^* and empty π^* orbitals (2) filled σ and filled π orbitals

(3) empty σ and empty π^* orbitals (4) empty σ^* and filled π orbitals

[ANS] **2**

[SOLN]



[Q.66] A 1:3 electrolyte an aqueous solution is

(1) $\left[Co(NH_3)_3(NO_2)_3\right]$ (2) $\left[CoCl_2(NH_3)_4\right]Cl$

(3) $\left[CoCl(NH_3)_5\right]Cl_2$ (4) $\left[Co(NH_3)_6\right]Cl_3$

[ANS] **4**

[Q.67] The standard electrode potential (E^o) for the half-cell reaction $Fe^{3+} + e^- \rightarrow Fe^{2+}$ at 298 K is

(Given: $E^o(Fe^{3+}/Fe) = -0.04\,V$ and $E^o(Fe^{2+}/Fe) = -0.44\,V$ at 298 K)

(1) $+0.92\,V$ (2) $+0.40\,V$ (3) $+0.76\,V$ (4) $-0.48\,V$

[ANS] **3**

[SOLN] $Fe^{3+} + e^- \rightarrow Fe^{+2}$

$Fe^{3+} + 3e^- \rightarrow Fe \quad E^0 = -0.04V \quad \Delta G^0 = -3F(-0.04)$

$Fe \rightarrow Fe^{+2} + 2e^- \quad E^0 = 0.44V \quad \Delta G^0 = -2F(0.44)$

$Fe^{3+} + e^- \rightarrow Fe^{+2} \quad E^0 = xV \quad \Delta G^0 = -1F(x)$

$-1F(x) = -3F(-0.04) - 2F(0.44)$

$x = 3(-0.04) + 2(0.44)$

$= -0.12 + 0.88$

$= 0.76V$

[Q.68] In potash alum, the ratio of K^+ and SO_4^{2-} ions is

(1) 3 : 2 (2) 1 : 2 (3) 2 : 1 (4) 2 : 3

[ANS] **2**

[Q.69] Consider the following statements about the solutions formed by mixing two liquids.

(A) An ideal solution thus formed obeys Raoult's law throughout the composition range.

(B) Mixture of chloroform and acetone shows negative deviation from Raoult's law.

(C) Mixture of aniline and phenol shows positive deviation from Raoult's law.

(1) A and C only (2) A and B only (3) B and C only (4) A only

[ANS] **2**

[SOLN] A and B only

[Q.70] For a salt XY, which is a strong electrolyte, the plot of Λ_m versus $\sqrt{c}$ has a slope of -90.0 S m^2 $mol^{-3/2}$ $L^{1/2}$ at 298 K. At 0.01 M concentration of XY, the value of Λ_m is 145.0 S cm^2 mol^{-1}. The limiting molar conductivity of Y^{-1} ions $\left(\lambda_Y^O -, \text{in S cm}^2\,\text{mol}^{-1}\right)$ at 298 K will be

(Given : $\lambda_X^O + = 74.0$ S cm^2 mol^{-1})

(1) 76.0 (2) 80.0 (3) 100.0 (4) 90.0

[ANS] **2**

[SOLN] XY

$\lambda_M = \lambda_M^0 - b\sqrt{c}$

$145 = \lambda_M^0 - 90\sqrt{10^{-2}}$

$145 = \lambda_M^0 - 9$

$\lambda_M^0 = 145 + 9 = 154$

$\lambda_M^0(XY) = \lambda_M^0(X^+) + \lambda_M^0(Y^-)$

$\lambda_M^0(Y^-) = 154 - 74 = 80$

[Q.71] Arrange the following compounds in the increasing order of polarity

A. $CH_3CH_2OCH_2CH_3$ B. CH_3CH_2OH

C. CH_3COCH_3 D. CH_3COCH

Choose the correct answer from the options given below.

(1) $A < C < B < D$ (2) $A < B < C < D$ (3) $C < A < D < B$ (4) $C < A < B < D$

[ANS] **1**

[SOLN] **D > B >C > A**

[Q.72] According to crystal field theory, the correct order of ligands with respect to their decreasing order of field strength is

(1) $Cl^- > NH_3 > H_2O > CO$ (2) $CO > NH_3 > H_2O > Cl^-$

(3) $CO > H_2O > NH_3 > Cl^-$ (4) $Cl^- > H_2O > NH_3 > CO$

[ANS] **2**

[SOLN] **Correct order of ligand strength**

$CO > NH_3 > H_2O > Cl^-$

[Q.73] The amino acid that gives a red-blood colour on treating its sodium fusion extract with sodium nitroprusside is

(1) Serine (2) Leucine (3) Threonine (4) Methionine

[ANS] **4**

[SOLN] Methionine

[Q.74] In an acidic medium, 10 mL of 0.25 M oxalic aid is titrated with $KMnO_4$ solution. If the volume of $KMnO_4$ solution required to reach end point is 10 mL, the strength of the $KMnO_4$ solution is

(1) 0.15 M (2) 0.10 M (3) 0.20 M (4) 0.25 M

[ANS] **2**

[SOLN] M.Eq. of $H_2C_2O_4$ = M.Eq. of $KMnO_4$

$$\left[MV\,(ml)\,n_f\right]_{H_2C_2O_4} = \left[MV\,(ml)\,n_f\right]_{KMnO_4}$$

$$0.25 \times 10 \times 2 = M \times 10 \times 5$$

M=0.1

[Q.75] The correct statement is

(1) Aluminium has five valence orbitals.

(2) Boron has a maximum covalency of four.

(3) Beryllium has three valence orbitals.

(4) Magnesium has a maximum covalency of four.

[ANS] **2**

[SOLN] $Al \rightarrow 3s^2 3p^1 \rightarrow 4$ valence orbitals

$B \rightarrow 2^{nd}$ period element $\rightarrow$ maximum covalency = 4

$Be \rightarrow 2s^2 2p^0 \rightarrow 4$ valence orbitals

$Mg \rightarrow 3^{rd}$ period element $\rightarrow$ max. covalency = 6

[Q.76] Among the following, the compound having conjugated double bonds is

(1) Hepta-1, 6-diene (2) Hepta-1, 3-diene (3) Hepta-1, 4-diene (4) Hepta-1, 5-diene

[ANS] **2**

[SOLN] hepta 1,3-diene

$C = C - C = C - C - C - C$

Conjugated diene

[Q.77] $2A \xrightarrow{k} B$ is a zero-order reaction, where k = 1.0 mol $L^{-1} min^{-1}$. If the initial concentration of A is 2 M, then the time taken to complete 75% of the reaction will be

(1) 2.0 min (2) 1.5 min (3) 0.75 min (4) 1.0 min

[ANS] **3**

[SOLN] $2A \xrightarrow{k} B$

$K = 1\ mol\,L^{-1}\ min^{-1}$

$K_A = 2K_r = 2$

$t_{75\%} = 1.5\ t_{50\%} = 1.5\ \dfrac{[A]_0}{2K_A}$

$t_{75\%} = 1.5 \times \dfrac{2}{2 \times 2}$

= 0.75 minute

[Q.78] The correct order of solubility of the given salts in water at 298 k is

Salt	K_{sp} at 298 K
AgBr	5.0×10^{-13}
$Zn(OH)_2$	1.0×10^{-15}
Hg_2Cl_2	1.3×10^{-18}

(1) $Zn(OH)_2 > AgBr > Hg_2Cl_2$ (2) $Hg_2Cl_2 > Zn(OH)_2 > AgBr$

(3) $AgBr > Zn(OH)_2 > Hg_2Cl_2$ (4) $Hg_2Cl_2 > AgBr > Zn(OH)_2$

[ANS] **1**

[SOLN] $AgBr \rightarrow s_1 = \sqrt{K_{sp(AgBr)}} = \sqrt{5x10^{-13}} = \sqrt{50}\ x10^{-7} = 0.707x10^{-6}\ M$

$Zn(OH)_2 \rightarrow s_2 = \sqrt[3]{\dfrac{K_{sp(Zn(OH)_2)}}{4}} = \sqrt[3]{\dfrac{10^{-15}}{4}} = \dfrac{1}{\sqrt[3]{4}}\ x10^{-5} = 0.629\ x10^{-5} = 6.29\ x10^{-6}\ M$

$Hg_2Cl_2 \rightarrow s_3 = \sqrt[3]{\dfrac{K_{sp(Hg_2Cl_2)}}{4}} = \sqrt[3]{\dfrac{1.3x10^{-18}}{4}} = \sqrt[3]{\dfrac{1.3}{4}}x10^{-6} = 0.687x10^{-6}\ M$

$s_2 > s_1 > s_3$

[Q.79] The correct decreasing order of oxidation state of underlined atom in each molecule is

(1) $\underline{P_4O_6} > \underline{Cl_2O_7} > \underline{Al}H_3$ (2) $\underline{P_4O_{10}} > \underline{S}O_3 > H_2\underline{O}$

(3) $\underline{N_2O_5} > Al_2O_3 > H_2\underline{S}$ (4) $\underline{Pb}O_2 > N_2O_3 > SO_3$

[ANS] **3**

[SOLN] **1.** $P_4O_6 \rightarrow 4P - 12 = 0$

$4P = 12 \Rightarrow P = +3$

$Cl_2O_7 \rightarrow 2Cl - 14 = 0 \Rightarrow Cl = +7$

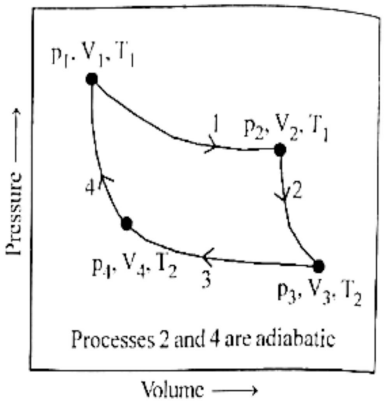
$AlH_3 \rightarrow Al - 3 = 0 \Rightarrow Al = +3$

2. $P_4O_{10} = P = +5, SO_3 \Rightarrow S = +6, H_2O \rightarrow O = -2$

3. $N_2O_5 \rightarrow N = +5,\ Al_2O_3, Al = +3,\ H_2S,\ S = -2$

4. $PbO_2 \rightarrow Pb = +4,\ N_2O_3 \rightarrow N = +3,\ SO_3, S = +6$

[Q.80] Consider the reversible process be 1.0 mol of an ideal gas as shown in the figure.



w1, w2, and w3 and w4 represent work done (in calories) in the processes 1, 2, 3 and 4, respectively, ΔU_2 and ΔU_4 are changes in the internal energy for the processes 2 and 4, respectively. [use R =2 cal K^{-1} mol^{-1}]

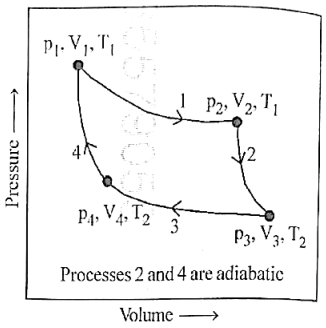
The correct option is

(1) $w_1 + w_2 + w_3 + w_4 = 0$ (2) $w_1 + w_3 = -2T_1 \ln \frac{V_2}{V_1} = 2T_2 \ln \frac{V_4}{V_3}$

(3) $w_1 + w_4 = \Delta U_2 - \Delta U_4$ (4) $w_1 + w_2 = 2T_1 \ln \frac{V_2}{V_1}$

[ANS] **2**

[SOLN]



$$W_1 = -1RT_1 \ln \frac{V_2}{V_1} = -2T_1 \ln \frac{V_2}{V_1}$$

$$W_3 = -1RT_2 \ln \frac{V_4}{V_3} = -2T_2 \ln \frac{V_4}{V_3}$$

$$W_1 + W_3 = -2T_1 \ln \frac{V_2}{V_1} - 2T_2 \ln \frac{V_4}{V_3}$$

[Q.81] **Assertion A :** For an ideal solution formed by mixing liquids P and Q, Δ_{mix} H = 0 and Δ_{mix} V = 0.

Reason R : No interactions occur between P and Q In the light of the above statements, choose the most appropriate answer from the options given below.

(1) A is not correct but R is correct.

(2) Both A and R are correct and R is the correct explanation of A.

(3) Both A and R are correct but R is NOT the correct explanation of A

(4) A is correct but R is not correct

[ANS] **4**

[SOLN] Assertion : Correct

Reason : Incorrect

[Q.82] Among the species given below, the spin-only magnetic moment is highest for

(Given : Atomic number Ti = 22, Mn = 25, Fe = 26 and Co = 27)

(1) $[Ti(H_2O)_6]^{3+}$ (2) $[Mn(CN)_6]^{3-}$ (3) $[Fe(CN)_6]^{3-}$ (4) $[Co(NH_3)_6]^{3+}$

[ANS] **2**

[SOLN] $\left[Mn(CN)_6\right]^{3-} \rightarrow Mn^{3+} \rightarrow 3d^4 \left(\text{low spin}\right) \rightarrow 2 \text{unpaired } e^-$

$\left[Ti(H_2O)_6\right]^{3T} \rightarrow Ti^{3+} \rightarrow 3d^1 \rightarrow 1$ unpaired electron

$\left[Fe(CN)_6\right]^{3-} \rightarrow Fe^{3+} \rightarrow 3d^5 \left(\text{low spin}\right) \rightarrow 1 \text{unpaired } e^-$

$\left[Co(NH_3)_6\right]^{3+} \rightarrow Co^{3+} \rightarrow 3d^6 \left(\text{low spin}\right) \rightarrow 0$ unpaired

[Q.83] A protein undergoes reversible thermal denaturation from its initial state N to denatured state D according to N $\rightleftharpoons$ D. At 60°C, the concentrations of both N and D are equal at equilibrium, and the standard enthalpy change of denaturation is 666 kJ mol–1. The standard entropy change $[\Delta S^\circ$ in kJ $K^{-1}mol^{-1}]$ of the protein upon denaturation at 60°C is closest to

(1) 11.1 (2) 2.0 (3) 2000.0 (4) 333.0

[ANS] **2**

$N \rightleftharpoons D.....K_{eq.} = 1.....as [N] = [D]$

[SOLN]

$\Delta H^0 = 666\,kJ\,mol^{-1}$

$\Delta G^0 = 0 = \Delta H^0 - T\Delta S^0$

$\Delta S^0 = \frac{\Delta H^0}{T} = \frac{666}{333} = 2\,kJ\,K^{-1}\,mol^{-1}$

[Q.84] Given below are two statements : One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Generally, 3d transition metals have high melting points.

Reasons R : Involvement of 3d-electrosn is addition to 4s-electrons in the interatomic metallic bonding.

In light of the above statements, choose the most appropriate answer from the options given below :

(1) A is not correct but R is correct.

(2) Both A and R are correct and R is the correct explanation of A.

(3) Both A and R are correct but R is NOT the correct explanation of A

(4) A is correct but R is not correct

[ANS] 2

[SOLN] High melting point of transition elements due to strong metallic bonding because 3d and 4s electron participate in metallic bonding.

[Q.85] Given below are two statements : One is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : The first ionization enthalpy of O is lower than that of N and F.

Reason R : The loss of an electron from O leads to stable half-filled p orbital.

In light of the above statements, choose the most appropriate answer from the options given below :

(1) A is not correct but R is correct.

(2) Both A and R are correct and R is the correct explanation of A.

(3) Both A and R are correct but R is NOT the correct explanation of A

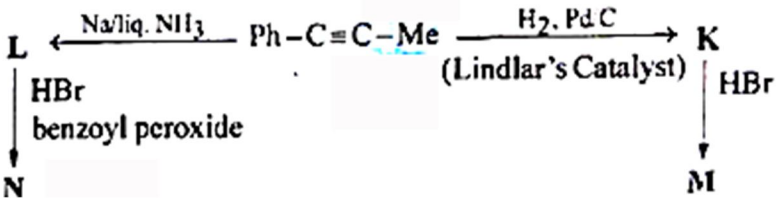
(4) A is correct but R is not correct

[ANS] 2

[SOLN] $IE_1 = F > N > O$

$O \xrightarrow{-e^-} O^+ \rightarrow 2p^3 \rightarrow$ stable configuration

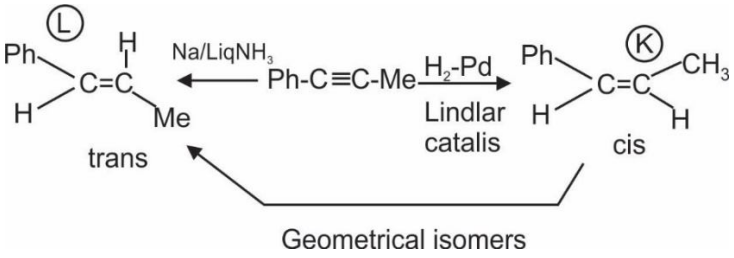
[Q.86] Consider the following reaction sequences and choose the correct option.



(1) M and N are stereoisomers
(2) K and L are geometrical isomers
(3) K and L are enantiomers
(4) M and N are geometrical isomers

[ANS] **2**

[SOLN]



[Q.87] The highest occupied molecular orbital for Ne_2 is

(1) σ^*_{2p} (2) π_{2p} (3) σ_{2p} (4) π^*_{2p}

[ANS] **1**

[SOLN] $Ne_2 \rightarrow 20$ electron $\rightarrow$ HOMO $= \sigma^*_{2p}$

[Q.88] Match the species in List I with their geometry in List II

List I	List II
A. PC_5	I. Tetrahedral
B. BrF_5	II. Square Planar
C. BrF_4^-	III. Trigonal bipyramidal
D. $[Ni(CN)_4]^{2-}$	IV. Square pyramidal

Choose the correct answer from the options given below:

(1) A – III, B – II, C – I, D - IV
(2) A – IV, B – III, C – I, D - II
(3) A – III, B – IV, C – I, D - II
(4) A – III, B – I, C – II, D - IV

[ANS] **3**

[SOLN] $PCl_5 \rightarrow$ trigonal bipyranidal

$BrF_5 \rightarrow$ Square pyramidal

BF_4^- → Tetrahedral

$\left[Ni(CN)_4\right]^{2-}$ → square planar

[Q.89] Match the vitamins in List I with their sources in List II

	List I		List II
A.	vitamin A	I.	Meat
B.	vitamin B_{12}	II.	sunflower oil
C.	vitamin E	III.	Green leafy vegetables
D.	vitamin K	IV.	carrots

Choose the correct answer from the options given below.

(1) A- lll, B-1, C-1V, D-II

(2) A-II, B-I, C-1V, D-I

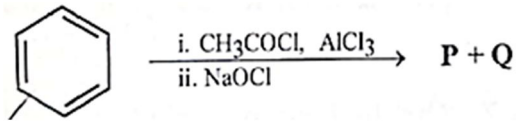
(3) A-1V, B-I, C-II, D-III

(4) A-IV, B-II, C-I, D-III

[ANS] **3**

[SOLN] **(NCERT)**

[Q.90] For the following reaction sequence, choose the correct option



(1) Both P and Q are carbonyl compounds.

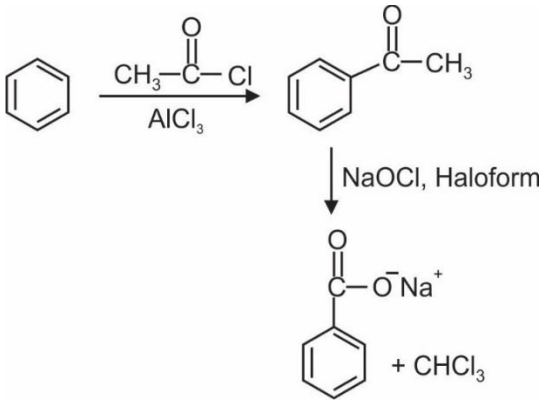
(2) If P is the sodium salt of a carboxylic, Q is a primary alcohol.

(3) P and Q are aromatic compounds.

(4) If P gives a carboxylic acid on acidification, Q gives a poisonous gas on exposure to air and light.

[ANS] **4**

[SOLN]



BIOLOGY

[Q.91] Given below are two statements :

Statement I : the class name Reptilia refers to creeping or crawling mode of locomotion.

Statement II : All organisms belonging to Reptilia have three chambered heart. In the light of the above statements, choose the **most appropriate** answer from the options given below :

(1) **Statement I** is incorrect but **Statement II** is correct

(2) Both **Statement I** and **Statement II** are correct

(3) Both **Statement I** and **Statement II** are incorrect

(4) **Statement I** is correct but **Statement II** is incorrect

[ANS] 4

[Q.92] How many turns of Calving cycle are required for the formation of three molecules of glucose?

(1) 18 (2) 6 (3) 3 (4) 1

[ANS] 1

[Q.93] Photorespiration reaction catalyzed by RuBisCo is shown below :

RuBP + $O_2 \rightarrow$ 3-Phosphoglycerate +X Identify "X" from given options :

(1) Malate (2) Phosphoenolpyruvate

(3) 2-Phosphoglycolate (4) Oxaloacetate

[ANS] 3

[Q.94] Given below are two statements :

Statement-I : in gymnosperms, the male and female gametophytes remain within the sporangia.

Statement-II : In gymnosperms, seeds are not covered.

In the light of the above statements, choose the **most appropriate** answer from the options given below:

(1) **Statement I** is incorrect but **Statement II** is correct

(2) Both **Statement I** and **Statement II** are correct

(3) Both **Statement I** and **Statement II** are incorrect

(4) **Statement I** is correct but **Statement II** is incorrect

[ANS] 2

[Q.95] How many molecules of pyruvic acid are produced at the end of glycolysis from 206 molecules of glucose ?

(1) 412 (2) 206 (3) 309 (4) 103

[ANS] 1

[Q.96] Match **List-I** with **List-II.**

	List-I		**List – II**
A.	Fusion of protoplasms between gametes	I.	Meiosis
B.	Fusion of two nuclei	II.	Plasmogamy
C.	generation of haploid spores	III.	Karyogamy

Choose the **correct** answer from the options given below :

(1) A-I, B-III, C-II (2) A-II, B-III, C-I (3) A-II, B-I, C-III (4) A-III, B-II, C-I

[ANS] 2

[Q.97] Mitochondrial inner membrane encloses _____.
(1) aqueous humor (2) matrix (3) cytosol (4) mucus
[ANS] **2**

[Q.98] Phyllotaxy is the pattern of arrangement of _____.
(1) sepals (2) leaves (3) flowers (4) fruits
[ANS] **2**

[Q.99] Mad cow disease is caused by _____.
(1) Mycoplasma sp. (2) prions (3) viroids (4) Aspergillus sp.
[ANS] **2**

[Q.100] Cell theory was formulated by _____.
(1) Antonie Von Leeuwenhoek (2) Schleiden and Schwann
(3) Robert Brown (4) Singer and Nicolson
[ANS] **2**

[Q.101] Which of the following plant growth regulators promotes internode elongation prior to flowering in cabbage ?
(1) Ethephon (2) Abscisic acid. (3) Gibberellin (4) Indole butyric acid
[ANS] **3**

[Q.102] Which pigment has absorption peak at 700 nm in the photosynthetic reaction centre PS I (P700)?
(1) Carotenoids (2) Chlorophyll b (3) Chlorophyll a (4) Xanthophylls
[ANS] **3**

[Q.103] Sphenopsida class belongs to_____.
(1) pteridophytes (2) bryophytes (3) angiosperms (4) gymnosperms
[ANS] **1**

[Q.104] Which of the following represents the correct sequence of arrangement of bones in the lower limb of humans ?
(1) Femur-tarsal-patella-tibia (2) Femur-tibia-patella-tarsal
(3) Patella-femur-tibia-tarsal (4) Femur-patella-tibia-tarsal
[ANS] **4**

[Q.105] Which of the following plant growth regulators is used as herbicide ?
(1) Gibberellin (2) 2,4-D (3) Kinetin (4) Abscisic acid
[ANS] **2**

[Q.106] Genus represents_____
(1) a group of closely related families (2) an individual plant or animal
(3) a population of plants and animals (4) a group of closely related species
[ANS] **4**

[Q.107] The plastid that stores xanthophyll is known as_____
(1) amyloplast (2) chloroplast (3) chromoplast (4) aleuroplast
[ANS] **3**

[Q.108] In water, frogs respire using_____
(1) trachea (2) skin (3) buccal cavity (4) lungs
[ANS] **2**

[Q.109] Which of the following is not a characteristic of chordates ?

(1) Presence of post anal part, (tail) (2) Presence of notochord,

(3) Central nervous system is dorsal (4) Absence of gills.

[ANS] 4

[Q.110] Smooth endoplasmic reticulum_____

(1) is a site for the synthesis of carbohydrates surface

(2) has ribosomes attached to its surface.

(3) is the major site for the synthesis of lipids

(4) is actively involved in protein synthesis

[ANS] 3

[Q.111] Which of the following are characteristics of prokaryotic cells?

(a) Ribosomes are made of 50S and 30S subunits

(b) They cany have plasmids

(c) They contain mesosome

(d) They have peroxisomes

Choose the correct answer from the options given below:

(1) (a), (b) and (c) only (2) (b) and (c) only

(3) (a) and (c) only (4) (a), (c) and (d) only.

[ANS] 1

[Q.112] Match List-I with List-II.

List-I	**List-II**
A. Cristate-	I. Flat membrane sacs I stroma of chloroplast
B. Cisternae	II. Infoldings in mitochondria
C. Thylakoids	III. Cell membrane
D. Phospholipid	IV. Disc shaped sacs in the Golgi apparatus

Choose the correct answer from the options given below :

(1) A – IV, B – III, C – I, D – II (2) A – III, B – IV, C – I, D – II

(3) A – II, B – IV, C – I, D – III (4) A – II, B – IV, C – III, D – I

[ANS] 3

[Q.113] Which of the following statement related to pituitary gland are correct

(a) It is divided anatomically into adenohypophysis neurohypophysis

(b) It secretes follicle stimulating hormone

(c) It secretes melanocyte stimulating hormone

(d) It does not secrete prolactin

Choose the correct answer from the options given below :

(1) (b) and (c) only (2) (a) and (b) only (3) (a), (b) and (c) only (4)(c) and (d) only

[ANS] 3

[Q.114] Which of the following statements regarding photorespiration are correct?

(a) Do not occur in C3 plants (b) CO_2 is consumed and O_2 is generated

(c) Phosphoglycolate is formed (d) No synthesis of ATP and NaDPH

Choose the correct answer from the options given below :

(1) (a) and (b) only (2) (a) and (b) only (3) (c) and (d) only (4) (b) and (d) only

[ANS] 3

[Q.115] Which of the following statements is incorrect?

(1) Fibrinogen is produced from fibrin (2) Blood coagulates in response to an injury

(3) Blood clot consists of fibrins (4) Fibrin is produced from fibrinogen

[ANS] 1

[Q.116] Arrange the following taxonomic categories in ascending order.

(a) Genus (b) Class (c) Order (d) Phylum

(e) Family (f) Kingdom (g) Species

Choose the correct answer from the options given below :

(1) (f), (c), (b), (g), (d), (e), (a) (2) (g), (a), (e), (c), (b), (d), (f)

(3) (a), (c), (d), (g), (f), (b), (e) (4) (g), (c), (d), (b), (e), (a), (f)

[ANS] 2

[Q.117] Select the correct sequence of experiment that led to a gradual understanding of photosynthesis in green plants.

(1) Production of glucose $\rightarrow$ role of air $\rightarrow$ release of oxygen $\rightarrow$ absorption spectra of chlorophyll a and b

(2) Absorption spectra of chlorophyll a and b $\rightarrow$ production of glucose $\rightarrow$ release of oxygen $\rightarrow$ role of air

(3) Role of air $\rightarrow$ release of oxygen $\rightarrow$ production of glucose $\rightarrow$ absorption spectra of chlorophyll a and b

(4) Release of oxygen $\rightarrow$ production of glucose $\rightarrow$ absorption spectra of chlorophyll a and b $\rightarrow$ role of air

[ANS] 3

[Q.118] Match List-I with List-II.

	List-I		**List-II**
A.	Starch	I.	Fights infection
B.	Antibody	II.	Energy storage
C.	Concanavalin	III.	Glucose transport
D.	Glut-4	IV.	Lectin

Choose the correct answer from the options given below:

(1) A – I, B – II, C – III, D – IV (2) A – I, B – II, C – IV, D – III

(3) A – II, B – I, C – IV, D – III (4) A – II, B – I, C – III, D – IV

[ANS] 3

[Q.119] The number of vertebrate in a human is _____.

(1) 206 (2) 7 (3) 12 (4) 26

[ANS] 4

[Q.120] Endomembrane system includes_____.

(1) Golgi complex, chloroplast, peroxisomes and vacuole

(2) endoplasmic reticulum, Golgi complex, lysosomes and vacuole

(3) endoplasmic reticulum, chloroplast, peroxisomes and vacuole

(4) mitochondria, chloroplast, peroxisome and vacuole

[ANS] 2

[Q.121] Length of the stem at time 0 is 20 cm. The arithmetic growth rate is 30 cm per day. What is the length of the stem at the end of the 7^{th} day?

(1) 460 cm (2) 50 cm (3) 170 cm (4) 230 cm

[ANS] 4

[Q.122] Match List-I with List-II.

	List-I		**List-II**
A.	Spherical	I.	Vibrio
B.	Rod	II.	Cocci
C.	Comma	III.	Spirilla
D.	Spirillum	IV.	Bacilli

Choose the correct answer from the option given below:

(1) A-II, B-IV, C-I, D-III (2) A-I, B-III, C-II, D-IV

(3) A-III, B-II, C-I, D-IV (4) A-II, B-I, C-IV, D-III

[ANS] 1

[Q.123] The number of action potentials generated by sino-arterial node (SAN) in a healthy human is _____ per minute.

(1) 120 – 140 (2) 28 – 30 (3) 70 – 75 (4) 100 – 110

[ANS] 3

[Q.124] Match List-I with List-II

	List – I		**List – II**
A.	Family	I.	Sapindales
B.	Genus	II.	Dicotyledonae
C.	Class	III.	Anacardiaceae
D.	Phylum	IV.	Angiospermae
E.	Order	V.	Mangifera

Choose the correct answer from the options given below:

(1) A-III, B-V, C-II, D-IV, E-I(2) A-I, B-V, C-II, D-IV, E-III

(3) A-II, B-I, C-III, D-IV, E-V(4) A-II, B-III, C-V, D-I, E-IV

[ANS] 1

[Q.125] Which of the following is not a part of human central neural system?

(1) Pericardium (2) Arachnoid (3) Dura mater (4) Pia mater

[ANS] 1

[Q.126] Given below are two statements :

Statement I : Chromosomes are fully condensed at the end of prophase I.

Statement II : Meiosis I resembles mitosis.

In the light of the above statements, choose the most appropriate answer from the options given below:

(1) Statement I is incorrect, but Statement II is true

(2) Both Statement I and Statement II are true

(3) Both Statement I and Statement II are false

(4) Statement I is correct, but Statement II is false

[ANS] 4

[Q.127] Match List-I with List-II.

	List-I		**List-II**
A.	Marginal placentation	I.	Argemone
B.	Axile placentation	II.	Tomato
C.	Parietal placentation	III.	Primrose
D.	Free central placentation	IV.	Pea

Choose the correct answer from the options given below:

(1) A-IV, B-II, C-I, D-III (2) A-II, B-IV, C-I, D-III

(3) A-IV, B-II, C-III, D-I (4) A-IV, B-III, C-I, D-II

[ANS] 1

[Q.128] Symbiotic association between fungi and algae are called _____.

(1) Chrysophytes (2) Lichens (3) Sponges (4) Mycorrhiza

[ANS] 2

[Q.129] Which of the following is not a prokaryote?

(1) Fungi (2) Bacteria (3) Blue green algae (4) Mycoplasma

[ANS] 1

[Q.130] Arrange the following elements in descending order of their contribution to percentage weight of the human body.

(a) Oxygen (b) Carbon (c) Hydrogen (d) Nitrogen

Choose the correct answer from the options given below:

(1) (b), (a), (c), (d) (2) (a), (b), (c), (d) (3) (c), (a), (b), (d) (4) (b), (c), (d), (a)

[ANS] 2

[Q.131] Which one of the following statements in incorrect?

(1) β-cells of pancreas secrete insulin (2) α-cells of pancreas secrete glucagon

(3) α-cells of pancreas secrete insulin (4) Glucagon stimulates glycogenolysis

[ANS] 3

[Q.132] Which of the following are characteristic features of Solanaceae family?

(a) Flowers are bisexual and actinomorphic

(b) Calyx have five sepals and are united

(c) Androecium have five stamens and are epipetalous

(d) Ovary is inferior

Choose the correct answer from the options given below:

(1) (b), (c) and (d) only (2) (a), (b) and (c) only

(3) (d) only (4) (a) and (b) only

[ANS] 2

[Q.133] Given below are two statements :

Statement I : When any plane passing through the central axis of the passing through the central axis of the body divides the organism into two identical halves, it is called radial symmetry.

Statement II : In phylum Echinodermata, both adults and larvae are radially symmetrical.

In the light of the above statements, choose the most appropriate answer from the options given below:

(1) Statement I is incorrect but Statement II is correct

(2) Both Statement I and Statement II are correct

(3) Both Statement I and Statement II are incorrect

(4) Statement I is correct but Statement II is incorrect

[ANS] 4

[Q.134] The correct sequence of adult cell cycle phases is _____.

(1) S-M-G2-G1 (2) G1-G2-S-M (3) G1-M-G2-S (4) G1-S-G2-M

[ANS] 4

[Q.135] In frogs, the number of pairs of cranial nerves arising from the brain are _____.

(1) 12 (2) 6 (3) 9 (4) 10

[ANS] 4

[Q.136] Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : In recombinant DNA technology, lysozyme is used for disrupting bacterial cells while cellulase is for plant cells.

Reason R : Isolation of genetic material needs disruption of cells.

In the light of the above statements, choose the appropriate answer form the options given below:

(1) A is not correct but R is correct

(2) Both A and R are correct and R is the correct explanation of A

(3) Both A and R are correct but R is not the correct explanation of A

(4) A is correct but R is not correct

[ANS] 2

[Q.137] The method of directly of injecting a sperm into ovum in assisted reproductive technology is called:

(1) Embryo transfer (ET) (2) Gamete intra fallopian transfer (GIFT)
(3) Zygote intra fallopian transfer (ZIFT) (4) Intra cytoplasmic sperm injection (ICSI)

[ANS] 4

[Q.138] Adaptive radiation in placental mammals and Australian Marsupials leading to similarity between distant species is an example of _____.

(1) Genetic drift (2) Divergent evolution
(3) Convergent evolution (4) Founder effect

[ANS] 3

[Q.139] Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : In an experiment, Mendel observed that the F1 progeny plants are all tall and none are dwarf.

Reason R : Stem height is a contrasting trait, which tall being dominant and dwarf being recessive.

In the light of the above statements, choose the most appropriate answer from the options given below:

(1) A is not correct but R is correct
(2) Both A and R are correct and R is the correct explanation of A
(3) Both A and R are correct but R is not the correct explanation of A
(4) A is correct but R is not correct

[ANS] 2

[Q.140] Arrange the following in descending order of number of species in the Amazonian rain forest.

(a) Plants (b) Birds (c) Fishes
(d) Invertebrates (e) Mammals

Choose the correct answer from the options given below:

(1) (b) > (a) > (d) > (c) > (e)(2) (c) > (b) > (d) > (e) > (a)
(3) (d) > (a) > (c) > (b) > (e)(4) (e) > (b) > (a) > (c) > (d)

[ANS] 3

[Q.141] Sponges exchange O_2 with CO_2 by_____.

(1) gills
(2) simple diffusion over their entire body surfaces
(3) moist cuticle
(4) tracheal tubes

[ANS] 2

[Q.142] For a person with blood group 'O', which of the following is **not** a possible combination of parents' blood group genotypes ?

(1) Father : I^AI^B and Mother : I^Ai (2) Father : I^Ai and Mother : I^Bi
(3) Father : I^Ai and Mother : I^Ai (4) Father : I^Bi and Mother : I^Bi

[ANS] 1

[Q.143] Given below are two Statement .

Statement I : Modern Homo sapiens arose in Australia and moved across continents.

Statement II : Homo sapiens arose around 75000 to 10000 years ago.

In the light of the abovc statements, choose thc **most appropriate** answer from thc options given below :

(1) **Statement I** is incorrect but **Statement II** is correct

(2) Both **Statement I** and **Statement II** are correct

(3) Both **Statement I** and **Statement II** are incorrect

(4) **Statement I** is correct **Statement II** is incorrect

[ANS] 1

[Q.144] Which of the following is used as an effective sedative and painkiller for treating post-surgery patients ?

(1) Anti-retroviral drugs (2) interferon

(3) Antibiotics (4) Morphine

[ANS] 4

[Q.145] Which of the following plant produces non-albuminous seeds ?

(1) Pea (2) Wheat (3) Maize (4) Barley

[ANS] 1

[Q.146] Given below are two statements : one is labelled as Assertion A and the other is labelled as Reason R.

Assertion A : Abingdon tortoise in Galapagos islands became extinct within a decade after goats were introduced.

Reason R : Goats were more efficient at browsing than Abingdon tortoise.

In the light of the above statements, choose the **most appropriate** answer from the options given below :

(1) **A** is not correct but **R** is correct

(2) Both **A** and **R** are correct and **R** is the correct explanation of **A**

(3) Both **A** and **R** are correct but **R** is not the correct explanation of **A**

(4) **A** is correct but **R** is not correct

[ANS] 2

[Q.147] Thc covering of ovum at ovulation is _____.

(1) chorion (2) endometrium (3) zona radiata (4) zona pellucida

[ANS] 4

[Q.148] Which of the following is used as a clot buster ?

(1) Statins (2) Streptokinase (3) Penicillin (4) Cyclosporin A

[ANS] 2

[Q.149] Which of the following structure is **not** a part of the malc reproductive system ?

(1) Infundibulum (2) Rete testis (3) Epididymis (4) Vasa efferentia

[ANS] 1

[Q.150] Given below are two statements

Statement I : Ovulation is caused by LH surge leading to rupture of Graafian follicles.

Statement II : Graafian follicle remaining after ovulation transform into corpus luteum and secretes large amount of estrogen.

In the light of the above statements, choose the **most appropriate** answer from the options given below

(1) **Statement I** is incorrect but **Statement II** is correct
(2) Both **Statement I** and **Statement II** are correct
(3) Both **Statement I** and **Statement II** are incorrect
(4) **Statement I** is correct but **Statement II** is incorrect

[ANS] 4

[Q.151] The opening between the right atrium and the right ventricle is guarded by ____.

(1) Sino-atrial node (2) bicuspid valve (3) tricuspid valve (4) Semilunar valve

[ANS] 3

[Q.152] Which of the following is not evidence for evolution?

(1) Divergent evolution of anatomical structures such as forelimbs
(2) Convergent evolution of traits like wings of birds and butterflies
(3) Paleontological evidence from fossil records
(4) Embryological support for evolution as proposed by Ernst Heckel

[ANS] 1

[Q.153] The inactive form of Bt toxin is converted to the active form in the insect gut _____

(1) by nucleases (2) due to alkaline pH
(3) due to acidic Ph (4) by proteases

[ANS] 2

[Q.154] Colostrum, secreted by mother during initial days of lactation, is abundant in_____.

(1) IgD (2) IgG (3) IgM (4) IgA

[ANS] 4

[Q.155] Which of the following in female gametophyte of an angiosperm helps in guiding the pollen tube for fertilizing the eggs?

(1) Polar nucleus (2) Antipodals (3) Synergids (4) Central cell

[ANS] 3

[Q.156] Which of the following disease is not sexually transmitted ?

(1) Genital warts (2) Syphilis (3) Tuberculosis (4) Gonorrhoea

[ANS] 3

[Q.157] Which of the following statements about lac-operon is correct?

(1) Galactose can act as an inducer of lac operon
(2) Gene i is constitutively expressed
(3) Lactose activates repressor to bind to the operator
(4) Genes i, z, y and a share single common Promoter.

[ANS] 2

[Q.158] Match List- I with List – II.

List – I	**List – II**
A. Transformation	I. Restriction enzyme
B. Cloning site	II. Transfer DNA to host bacteria
C. Selection	III. Replication
D. Ori	IV. Antibiotic

Choose the correct answer from the options given below:

(1) A –IV, B – I, C–III, D–II (2) A –II, B – I, C–IV, D–III

(3) A –I, B – II, C–IV, D–III (4) A –III, B – IV, C–II, D–I

[ANS] 2

[Q.159] A population of diploid organisms is at Hardy-Weinberg equilibrium. If the frequency of allele A is 0.1, the frequency of AA is___

(1) 0.99 (2) 0.01 (3) 0.02 (4) 0.10

[ANS] 2

[Q.160] Sperm motility is due to___

(1) Muscular Movement (2) Flagellar Movement

(3) Ciliary Movement (4) Amoeboid Movement

[ANS] 2

[Q.161] Consider a population of 10 million cells. Given the per-capita birth rate of 0.002 (per unit time) and the per-capita death rate of 0.002 (per unit time), the expected number of cells after 10 generations is

(1) 100 million (2) 1 million (3) 5 million (4) 10 million

[ANS] 4

[Q.162] Given below are two statements one, is labelled as **Assertion A** and the other is labelled as **Reason R**.

Assertion A: Forelimbs of human and bats are homologous.

Reason R: Forelimbs of humans, and bats have similar anatomical structure.

In the light of the above statements, choose the most appropriate answer from the options given below:

(1) A is false but R is true

(2) Both A and R are correct and R is the correct explanation of A

(3) Both-A and R are true, but R is not the correct explanation of A

(4) A is true but R is false

[ANS] 2

[Q.163] Muscle contraction is initiated by a signal sent by the central nervous system by the release of _____.

(1) cyclic adenine monophosphate (2) acetyl choline

(3) acetyl coenzyme A (4) cyclic guanine monophosphate

[ANS] 2

[Q.164] Which of the following hormone is not secreted by human placenta?

(1) LH (2) hCG (3) Estrogen (4) Progesterone

[ANS] 1

[Q.165] Which of the following statements is correct about *Plasmodium*?

(1) Fertilization takes place in mosquito gut

(2) Reproduces sexually in liver cells

(3) Reproduces sexually in RBCs

(4) Gametocytes develop in mosquito gut

[ANS] 1

[Q.166] Which of the following are primary consumers in a food chain?

(1) Carnivores (2) Parasites (3) Predators (4) Herbivores

[ANS] 4

[Q.167] Which of the following statements about the reabsorption process in Henle's loop arc correct?

(a) The descending limb of Henle's loop is permeable to water but almost impermeable to electrolytes.

(b) Urine gets concentrated in Henle's loop.

(c) Reabsorption of Na^+ and water takes place in Henle's loop.

(d) Active or passive transport of electrolytes occurs in the ascending limb of Henle's loop.

Choose the correct answer from the options given below:

(1) (a), (b) and (d) only (2) (a) and (b) only

(3) (b), (c) and (d) only (4) (a), (b) and (c) only

[ANS] 1

[Q.168] Given below are two statements: one is labelled as **Assertion A** and the other is labelled as **Reason R**.

Assertion A: The logistic growth model of populations is considered more realistic than the exponential growth model.

Reason R: Resources are finite.

In the light of the above statements,

Choose the most appropriate answer from the options given below:

(1) A is not correct but R is correct

(2) Both A and R are correct and R is the correct explanation of A

(3) Both A and R are correct but R is not the correct explanation of A

(4) A is correct but R is not correct

[ANS] 2

[Q.169] Which of the following is the correct order of arrangement of vertebrate column from the head to toe?

(1) Cervical vertebra, thoracic vertebra, lumbar vertebra, sacrum

(2) Cervical vertebra, thoracic vertebra, sacrum, lumbar vertebra

(3) Sacrum, lumbar vertebra, thoracic vertebra, cervical vertebra

(4) Cervical vertebra, lumbar vertebra, thoracic vertebra, sacrum

[ANS] 1

[Q.170] Match **List – I** with **List – II.**

List – I	**List – II**
A. Both species are harmed	I. Predation
B. One species is harmed and the other is benefited	II. Mutualism
C. Both species are benefited	III. Competition
D. One is benefited while the other has No effect	IV. Commensalism

Choose the correct answer from the options given below:

(1) A-III, B-1, C-II, D-IV (2) A-II, B-IV, C-II, D-I

(3) A-II, B-IV, C-II, D-I (4) A-I, B-I, C-1V, D-III

[ANS] 1

[Q.171] If the diploid chromosome number of typical angiosperm is 36, what would be the chromosome number in its endosperm ?

(1) 72 (2) 18 (3) 36 (4) 54

[ANS] 4

[Q.172] Which of the following enzymes synthesizes precursor mRNA ?

(1) DNA polymerase (2) RNA polymerase I

(3) RNA polymerase II (4) RNA polymerase III

[ANS] 3

[Q.173] Given below are two statements :

Statement I :

Plasmids are autonomously replicating DNA.

Statement II : Plasmids are extrachromosomal DNA. In the light of the above statements, choose the most appropriate answer from the options given below :

(1) Statement I is incorrect but Statement II is correct

(2) Both Statement I and Statement II are correct

(3) Both Statement I and Statement II are incorrect

(4) Statement I is correct but Statement II is incorrect

[ANS] 2

[Q.174] How many theca are present in each lobe of a typical bilobed angiosperm anther ?

(1) 12 (2) 2 (3) 6 (4) 8

[ANS] 2

[Q.175] Natural selection can lead to _____ .

(a) stabilisation (b) genetic drift

(c) directional change (d) disruption

Choose the correct answer from the options given below :

(1) (a) and (c) only (2) (a) only

(3) (a), (c) and (d) only (4) (a), (b), (c) and (d)

[ANS] 3

[Q.176] Which of the following statements are correct ?
(a) Energy flow from producers to consumers is unidirectional
(b) Energy pyramid can never be inverted
(c) Transfer of energy follows the 1% law
Choose the correct answer from the options given below :
(1) (b) and (c) only (2) (a), (b) and (c) (3) (a) and (b) only (4) (a) and (c) only

[ANS] 3

[Q.177] Match **List-I** with **List-II**.

	List-I		**List-II**
A.	Excess growth hormone	I.	Reabsorption of water and electrolytes in kidney
B.	Luteinizing hormone	II.	Contraction of uterus during child birth
C.	Vasopressin	III.	Acromegaly
D.	Oxytocin	IV.	Ovulation

Choose the correct answer from the options given below:
(1) A-IV, B-III, C-I, D-II (2) A-III, B-IV, C-II, D-I
(3) A-III, B-IV, C-I, D-II (4) A-II, B-IV, C-I, D-III

[ANS] 3

[Q.178] Which of the following are secondary lymphoid organs ?
(a) Bone marrow (b) Tonsils (c) Spleen (d) Thymus
Choose the correct answer from the options given below :
(1) (a) and (d) only (2) (a) and (b) only
(3) (b) and (c) only (4) (b) and (d) only

[ANS] 3

[Q.179] During PCR, primers bind to the DNA strands in the _____ step.
(1) ligation (2) denaturation
(3) extension (4) annealing

[ANS] 4

[Q.180] Given below are two statements :
Statement I : Down's syndrome is caused by the absence of one of the X-chromosomes.
Statement II : Turner's syndrome is caused by the presence of an additional copy of the chromosomes.
In the light of the above statements, choose the correct answer from the options given below :
(1) Statement I is incorrect but Statement II is correct
(2) Both Statement I and Statement II are correct
(3) Both Statement I and Statement II are incorrect
(4) Statement I is correct but Statement II is incorrect

[ANS] 3